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Student Exploration: Moles

Vocabulary: atomic mass, Avogadro constant, conversion factor, dimensional analysis, mole, molar mass, molecular mass, scientific notation, significant figures, unified atomic mass unit

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. In the image to the right, note a dozen eggs, a dozen donuts and a dozen roses. How many of each item do you have? 12

2. Would a dozen of each object have the same mass? NO

3. Suppose you have a dozen carbon atoms, a dozen gold atoms, and a dozen iron atoms. Even though you have the same number of each, would you expect them all to have the same mass? Explain.

No I would not expect them to have the same mass because they are made different atoms.



Gizmo Warm-up

When counting roses, eggs, or donuts, a dozen is a good unit to use. If you are counting atoms, however, a dozen is not much help. In the Moles Gizmo, you will learn about a unit used to count atoms.

On the AVOGADRO CONSTANT tab, place the copper (Cu) atom on the nano-balance on the left, which will show the average atomic mass of copper rather than the mass of a single copper atom.



1. What is the average mass of a copper atom? 63.546 u

The unit "u" refers to **unified atomic mass units**. A single proton or neutron has a mass of approximately one atomic mass unit. (Officially, 1 u is one-twelfth the mass of a C-12 atom.)

2. To gain an idea as to how many atoms are in a gram or so of copper, use the larger balance on the right. Press **Add atoms** to put a scoop of atoms in the weighing dish, and keep adding until the balance registers between 1 and 2 grams. If you don't seem to be making much progress, adjust the exponent using the slider, which will make the scoop size bigger.

How many atoms did you need to add? 1.01×10^{20}

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