

Name _____

Section _____

Experiment 10

Advance Study Assignment: Analysis of an Aluminum-Zinc Alloy

1. On the following page, construct a graph of N_{H_2} vs. % Al. To do this, refer to Equation 11 and the discussion preceding it. Note that a plot of N_{H_2} vs. % Al should be a straight line (why?). To fix the position of a straight line, it is necessary to locate only two points. The most obvious way to do this is to find N_{H_2} when % Al = 0 and when % Al = 100. If you wish you may calculate some intermediate points (for example, N_{H_2} when % Al = 50, or 20, or 80) all these points should lie on the same straight line. To use a spreadsheet, set up Equation 11 for different Al percentages. Graph N_{H_2} vs. % Al.
2. A student obtained the following data in this experiment. Fill in the blanks in the data and make the indicated calculations:

Mass of gelatin capsule	0.1134 g	Ambient temperature, t	23°C
Mass of capsule plus alloy sample	0.3212 g	Ambient temperature, T	296 K
Mass of alloy sample, m	0.2084 g	Atmospheric pressure	732 mm Hg
Mass of empty beaker	154.3 g	Vapor pressure of H_2O at t (Appendix I)	21.1 mm Hg
Mass of beaker plus displaced water	401.6 g		$732 - 21.1 = 710.9 \text{ mm Hg}$
Mass of displaced water (density = 1.00 g/mL)	247.3 g	Pressure of dry H_2 , P_{H_2} (Eq. 7)	710.7 mm Hg
Volume of displaced water	247.3 mL	Pressure of dry H_2	0.9354 atm
Volume, V , of H_2 = Volume of displaced water	247.3 mL		0.2473 L

Find the number of moles of H_2 evolved, n_{H_2} , (Eq. 5; V in liters, P_{H_2} in atm, T in K, $R = 0.0821$ liter-atm/mole K).

$$0.935 \times 0.2473 \times 296 = 0.0095 \text{ moles } H_2$$

Find, N_{H_2} , the number of moles of H_2 per gram of sample (n_{H_2}/m).

$$\frac{0.0095}{0.2084} = 0.0456 \text{ moles } H_2/g$$

Find the % Al in the sample from the graph prepared for Problem 1.

Find the % Al in the sample by using Equation 11.

_____ % Al

_____ % Al006C

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SA Dillow



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