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### Stoichiometry and Percent Yield Practice

Answer the following questions about theoretical yield and percent yield. Make sure you show all work, units, and round for sig figs.

1. Use the following equation to answer the questions below.



If I start this reaction with 57.0 grams of nitrogen with excess hydrogen, how many grams of ammonia ( $\text{NH}_3$ ) can theoretically be produced?

$$57.0 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28.014 \text{ g N}_2} \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \times \frac{17.031 \text{ g NH}_3}{1 \text{ mol NH}_3} = \boxed{69.3 \text{ g NH}_3}$$

If 60.0 grams of ammonia is actually produced, what was my percent yield of ammonia?

$$\frac{60.0}{69.3} \times 100 = \boxed{86.6\%}$$

2. Use the following equation to answer the questions below.



If 11.3 grams of carbon dioxide gas is used in this reaction, how many grams of lithium carbonate can theoretically be produced?

$$11.3 \text{ g CO}_2 \times \frac{1 \text{ mol CO}_2}{44.009 \text{ g CO}_2} \times \frac{1 \text{ mol Li}_2\text{CO}_3}{1 \text{ mol CO}_2} \times \frac{73.87 \text{ g Li}_2\text{CO}_3}{1 \text{ mol Li}_2\text{CO}_3} = \boxed{19.0 \text{ g Li}_2\text{CO}_3}$$

If 14.4 grams of lithium carbonate are actually produced, what was my percent yield?

$$\frac{14.4}{19.0} \times 100 = \boxed{75.8\%}$$

# Experiment 7 Reaction Stoichiometry Percent Yield Answers

**John A. Olmsted, Gregory M. Williams**



## **Experiment 7 Reaction Stoichiometry Percent Yield Answers:**

*Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

**SAT Subject Test Chemistry** Kaplan Test Prep, 2017-01-03 Kaplan's SAT Subject Test Chemistry is the most up-to-date guide on the market with the essential content, practice, and strategies students need for success on Test Day. Kaplan's expert tips and focused review will help you ace the test and give your college applications a boost. Essential Review: Three full-length practice tests with detailed answer explanations. A full-length diagnostic test identifies areas for score improvement so you can personalize your prep. Focused chapter summaries, highlights, and quizzes. End-of-chapter quizzes for additional practice. Proven score-raising strategies teach you how to tackle the test efficiently. Expert Guidance: We know the test. Our Learning Engineers have put tens of thousands of hours into studying the SAT, using real data to design the most effective strategies and study plans. Kaplan's expert psychometricians make sure our practice questions and study materials are true to the test. We invented test prep. Kaplan www.kaptest.com has been helping students for almost 80 years, and more than 95% of our students get into their top-choice schools. Our proven strategies have helped legions of students achieve their dreams.

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