

- 2) Light the Bunsen burner and place it under the beaker. Adjust the burner so the hottest part of the flame is touching the bottom of the beaker.
- 3) Carefully heat and stir the mixture in the beaker using your stir rod. The solution should be hot, but not boiling. After all the crystals have dissolved, remove the beaker from the heat using your hot pad or rubber-coated tongs.
- 4) Add 2.00g of iron filings slowly to the hot copper (II) sulfate solution while stirring. Record observations in your data table.
- 5) Allow the beaker to cool for 10-15 minutes.
- 6) Pour off (decant) the solution into a different beaker. Pouring the solution down a stirring rod is recommended (as shown in the image shown to the right). Make sure not to disturb the copper in the beaker.
- 10) Add a small amount of water (10-15mL) to the copper & stir.
- 11) Let the copper settle and decant again.
- 12) Dry the copper as your teacher directs and determine its mass. Record the mass in the data table provided.



Data Table:

Mass of clean, dry 250mL beaker		g
Observations after adding iron filings		
	Circle the observations that support the chemical reaction.	
Mass of beaker with dried copper residue		g
Mass of copper residue		g

Analysis Questions:

- 1) Which reactant is limiting? What observations support this? Use the balanced chemical equation for the reaction that occurred in the lab and stoichiometric calculations to support your answer.
- 2) Why was the copper washed with water in step 10 of the procedure?
- 3) Why didn't the water added to the copper (II) sulfate need to be measured exactly?
- 4) What is the formula for the iron-containing compound that is formed when copper (II) sulfate and iron react? Support your answer.

Cleaning Up:

- 1) Place the copper in the waste container provided by your teacher.
- 2) Place the remaining solution in the labeled container.
- 3) Clean up all materials and return them to the proper locations.
- 4) Wash your hands thoroughly before leaving the lab.

Conclusions: Provide a summary of the lab and discuss a final confirmation or rejection of your hypothesis, citing your observations.

Experiment 37 Stoichiometry Answers

JE Gale



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