

3. pH of buffer in which $[\text{HB}]/[\text{B}^-] = 0.10$ _____

pK_a _____

pH after addition of excess NaOH _____

Explain your observations.

4. pH of distilled water _____

pH after addition of 5 drops HCl _____

pH after addition of 5 drops NaOH _____

Explain your observations.

5. pH of buffer solution to be prepared _____

Mean value of pK_a (as found in Parts 1, 2, and 3) _____

$[\text{B}^-]/[\text{HB}]$ in buffer _____

Volume 0.10 M NaB/Volume 0.10 M HB needed in buffer _____

Volume 0.10 M NaB used _____ mL

Volume 0.10 M HB used _____ mL

pH of buffer you prepared _____

D. Preparing a Buffer from a Solution of a Weak Acid

pH of buffer to be prepared _____ pK_a of acid _____

Ratio of $[\text{B}^-]/[\text{HB}]$ required in buffer _____

Volume of 0.10 M NaOH calculated _____ mL

pH of acid solution before titration _____

Initial volume reading of NaOH _____ mL

Final volume reading of NaOH _____ mL

Volume of NaOH actually required _____ mL

How does that volume compare to your calculated value?

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DP Hallahan



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