

Experiment 23

Data and Calculations: Determination of the Equilibrium Constant for a Chemical Reaction

Mixture	Volume in mL, 2.00×10^{-3} M $\text{Fe}(\text{NO}_3)_3$	Volume in mL, 2.00×10^{-3} M KSCN	Volume in mL, Water	Method I Absorbance	Method II Depth in mm		$[\text{FeSCN}^{2+}]$
					Standard	Unknown	
1	5.00	1.00	<u>4.00</u>	<u>0.084</u>	—	—	<u>0.25×10^{-4} M</u>
2	5.00	2.00	<u>3.00</u>	<u>0.218</u>	—	—	<u>0.45×10^{-4} M</u>
3	5.00	3.00	<u>2.00</u>	<u>0.317</u>	—	—	<u>0.60×10^{-4} M</u>
4	5.00	4.00	<u>1.00</u>	<u>0.447</u>	—	—	<u>0.84×10^{-4} M</u>
5	5.00	5.00	<u>0</u>	<u>0.514</u>	—	—	<u>1.00×10^{-4} M</u>

If Method II was used, $[\text{FeSCN}^{2+}]_{\text{unknown}} = \text{—} \times 10^{-4}$ M; $[\text{FeSCN}^{2+}]$ in Mixtures 1 to 5 is found by Equation 7.

Processing the Data

A. Calculation of K_c assuming the reaction:



This calculation is most easily done by following Steps 1 through 5 in the discussion. Results are to be entered in the table on the following page. If you are using Excel, set up the table as we have, and follow the directions below.

Step 1 Find the initial number of moles of Fe^{3+} and SCN^{-} in the mixtures in test tubes 1 through 5. Use Equation 3 and enter the values in the first two columns of the table.

Step 2 Enter the experimentally determined value of $[\text{FeSCN}^{2+}]$ at equilibrium for each of the mixtures in the next to last column in the table. Use Equation 3 to find the number of moles of FeSCN^{2+} in each of the mixtures, and enter the values in the fifth column of the table. Note that this is also the number of moles of Fe^{3+} and SCN^{-} that were used up in the reaction.

Experiment 23 Determination Equilibrium Constant

Jerry L. Mills, Roy E. Mitchell



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Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations **General Chemistry**

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Moody, 2017-06-09 The definitive guide to the principles and practice of experimental organic chemistry fully updated and now featuring more than 100 experiments The latest edition of this popular guide to experimental organic chemistry takes students from their first day in the laboratory right through to complex research procedures All sections have been updated to reflect new techniques equipment and technologies and the text has been revised with an even sharper focus on practical skills and procedures The first half of the book is devoted to safe laboratory practice as well as purification and analytical techniques particularly spectroscopic analysis The second half contains step by step experimental procedures each one illustrating a basic principle or important reaction type Tried and tested over almost three decades over 100 validated experiments are graded according to their complexity and all are chosen to highlight important chemical transformations and to teach key experimental skills New sections cover updated health and safety guidelines additional spectroscopic techniques electronic notebooks and record keeping and techniques such as semi automated chromatography and enabling technologies such as the use of microwave and flow chemistry New experiments include transition metal catalysed cross coupling organocatalysis asymmetric synthesis flow chemistry and microwave assisted synthesis Key aspects of this third edition include Detailed descriptions of the correct use of common apparatus used in the organic laboratory Outlines of practical skills that all chemistry students must learn Highlights of aspects of health and safety in the laboratory both in the first section and throughout the experimental procedures Four new sections reflecting advances in techniques and technologies from electronic databases and information retrieval to semi automated chromatography More than 100 validated experiments of graded complexity from introductory to research level A user friendly experiment directory An instructor manual and PowerPoint slides of the figures in the book available on a companion website A comprehensive guide to

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