

2. **Optional** Assume that the reaction studied in Problem 1 is $\text{Fe}^{3+}(\text{aq}) + 2 \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{SCN})_2^{+}(\text{aq})$. Find K_c for this reaction, given the data in Problem 1.

a. Formulate the expression for K_c for the alternate reaction just cited.

$$K_c = \underline{\hspace{2cm}}$$

b. Find K_c as you did in Problem 1; take due account of the fact that two moles SCN^{-} are used up per mole $\text{Fe}(\text{SCN})_2^{+}$ formed as you carry out the following calculations:

Step 1 Results are as in Problem 1.

Step 2 How many moles of $\text{Fe}(\text{SCN})_2^{+}$ are in the mixture at equilibrium? (This will be the same as in Problem 1.)

$$\underline{\hspace{2cm}} \text{ moles } \text{Fe}(\text{SCN})_2^{+}$$

How many moles of Fe^{3+} and SCN^{-} are used up in making the $\text{Fe}(\text{SCN})_2^{+}$?

$$\underline{\hspace{2cm}} \text{ moles } \text{Fe}^{3+}; \underline{\hspace{2cm}} \text{ moles } \text{SCN}^{-}$$

Step 3 How many moles of Fe^{3+} and SCN^{-} remain in solution at equilibrium? Use the results of Steps 1 and 2, noting that moles SCN^{-} at equilibrium = original moles SCN^{-} - (2 $\times$ moles $\text{Fe}(\text{SCN})_2^{+}$).

$$\underline{\hspace{2cm}} \text{ moles } \text{Fe}^{3+}; \underline{\hspace{2cm}} \text{ moles } \text{SCN}^{-}$$

Step 4 What are the concentrations of Fe^{3+} , SCN^{-} , and $\text{Fe}(\text{SCN})_2^{+}$ at equilibrium? (Use Eq. 3 and the results of Step 3.)

$$[\text{Fe}^{3+}] = \underline{\hspace{2cm}} \text{ M}; [\text{SCN}^{-}] = \underline{\hspace{2cm}} \text{ M}; [\text{Fe}(\text{SCN})_2^{+}] = \underline{\hspace{2cm}} \text{ M}$$

Step 5 Calculate K_c based on the assumption that the alternate reaction occurs. (Use the answer to Part 2a.)

$$K_c = \underline{\hspace{2cm}}$$

Experiment 23 Determination Equilibrium Constant Answers

Thomas Griffiths



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Table of Contents Experiment 23 Determination Equilibrium Constant Answers

1. Understanding the eBook Experiment 23 Determination Equilibrium Constant Answers
 - The Rise of Digital Reading Experiment 23 Determination Equilibrium Constant Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Experiment 23 Determination Equilibrium Constant Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Experiment 23 Determination Equilibrium Constant Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Experiment 23 Determination Equilibrium Constant Answers
 - Personalized Recommendations

- Experiment 23 Determination Equilibrium Constant Answers User Reviews and Ratings
- Experiment 23 Determination Equilibrium Constant Answers and Bestseller Lists
- 5. Accessing Experiment 23 Determination Equilibrium Constant Answers Free and Paid eBooks
 - Experiment 23 Determination Equilibrium Constant Answers Public Domain eBooks
 - Experiment 23 Determination Equilibrium Constant Answers eBook Subscription Services
 - Experiment 23 Determination Equilibrium Constant Answers Budget-Friendly Options
- 6. Navigating Experiment 23 Determination Equilibrium Constant Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Experiment 23 Determination Equilibrium Constant Answers Compatibility with Devices
 - Experiment 23 Determination Equilibrium Constant Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Experiment 23 Determination Equilibrium Constant Answers
 - Highlighting and Note-Taking Experiment 23 Determination Equilibrium Constant Answers
 - Interactive Elements Experiment 23 Determination Equilibrium Constant Answers
- 8. Staying Engaged with Experiment 23 Determination Equilibrium Constant Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Experiment 23 Determination Equilibrium Constant Answers
- 9. Balancing eBooks and Physical Books Experiment 23 Determination Equilibrium Constant Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Experiment 23 Determination Equilibrium Constant Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Experiment 23 Determination Equilibrium Constant Answers
 - Setting Reading Goals Experiment 23 Determination Equilibrium Constant Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Experiment 23 Determination Equilibrium Constant Answers
 - Fact-Checking eBook Content of Experiment 23 Determination Equilibrium Constant Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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