

10.3**THE NATURE OF SOLIDS****SECTION REVIEW****Objectives**

- Describe how the degree of organization of particles distinguishes solids from gases and liquids
- Distinguish between a crystal lattice and a unit cell
- Explain how allotropes of an element differ

Key Terms

- melting point
- crystal
- unit cell
- allotropes
- amorphous solids
- glasses

Part A Completion

Use your completion material to check your understanding of the concepts and terms discussed previously in this section. Each blank uses the completed word or phrase, shown please, or name here.

Solids tend to be dense and _____. They do not flow or take the shape of their containers, like liquids do, because the particles in solids vibrate around _____ points. When a solid is sufficiently heated and its particles vibrate so strongly that they are no longer held in fixed positions, the solid _____. The _____ is the temperature at which a solid turns to liquid. The melting and _____ of a substance are at the same temperature. In general, ionic solids tend to have relatively _____ melting points while molecular solids tend to have relatively low melting points. Most solids are _____ in nature. The particles are arranged in a pattern known as a crystal _____. The smallest unit of a crystal lattice is the _____. Some solids have a disordered array of particles and are _____.

- _____
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