

PHYSICAL: Kinetic Molecular Theory

Learning Objectives

- At standard atmospheric conditions, all gases have approximately the same kinetic energy, and are made up of molecules that behave as independent particles (the particles of the gases are relatively small).
- At standard atmospheric conditions, all gases have the same average kinetic energy.
- At standard atmospheric conditions, all gases have the same average molecular mass.

Model 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in three different phases of matter.



Figure 1: Representation of particles in different phases of matter.

At standard atmospheric conditions, all gases have the same average kinetic energy, and are made up of molecules that behave as independent particles (the particles of the gases are relatively small).

Key Concepts

- At standard atmospheric conditions, all gases have the same average kinetic energy, and are made up of molecules that behave as independent particles (the particles of the gases are relatively small).
- At standard atmospheric conditions, all gases have the same average kinetic energy, and are made up of molecules that behave as independent particles (the particles of the gases are relatively small).

Figure 2: Representation of particles in different phases of matter.

	GAS	LIQUID	CRYSTALLINE SOLID
How are the particles arranged?	The particles are arranged in a regular, repeating pattern. The particles are relatively small and move randomly.	The particles are arranged in a regular, repeating pattern. The particles are relatively small and move randomly.	The particles are arranged in a regular, repeating pattern. The particles are relatively small and move randomly.
How are the particles arranged?	The particles are arranged in a regular, repeating pattern. The particles are relatively small and move randomly.	The particles are arranged in a regular, repeating pattern. The particles are relatively small and move randomly.	The particles are arranged in a regular, repeating pattern. The particles are relatively small and move randomly.

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The Molecular Theory of Solutions Ilya Prigogine, 1957 The Potential Distribution Theorem and Models of Molecular Solutions Tom L. Beck, Michael E. Paulaitis, Lawrence R. Pratt, 2006-08-31 An understanding of statistical thermodynamic molecular theory is fundamental to the appreciation of molecular solutions This complex subject has been simplified by the authors with down to earth presentations of molecular theory Using the potential distribution theorem PDT as the basis the text provides a discussion of practical theories in conjunction with simulation results The authors discuss the field in a concise and simple manner illustrating the text with useful models of solution thermodynamics and numerous exercises Modern quasi chemical theories that permit statistical thermodynamic properties to be studied on the basis of electronic structure calculations are given extended development as is the testing of those theoretical results with ab initio molecular dynamics simulations The book is intended for students taking up research problems of molecular science in chemistry chemical engineering biochemistry pharmaceutical chemistry nanotechnology and biotechnology *Kinetic Molecular Theory* Charles Proteus Steinmetz, American Electrochemical Society. Meeting, 1908* **Molecular Theory of Water and Aqueous Solutions** Arieh Ben-Naim, 2009 **Molecular Theory of Water and Aqueous Solutions** Arieh Ben-Naim, 2011 The aim of this book is to explain the unusual properties of both pure liquid water and simple aqueous solutions in terms of the properties of single molecules and interactions among small numbers of water molecules It is mostly the result of the author's own research spanning over 40 years in the field of aqueous solutions Jacket *Foundations of the Molecular Theory* John Dalton, Amedeo Avogadro, Joseph Louis Gay-Lussac, 1911 **Molecular Theory of Solutions** Arieh Ben-Naim, 2006-07-27 This book presents new and updated developments in the molecular theory of mixtures and solutions It is based on the theory of Kirkwood and Buff which was published more than fifty years ago This theory has been dormant for almost two decades It has recently become a very powerful and general tool to analyze study and understand any type of mixtures from the molecular or the microscopic point of view The traditional approach to mixture has been for many years based on the study of excess thermodynamic quantities This provides a kind of global information on the system The new approach provides information on the local properties of the same system Thus the new approach supplements and enriches our information on mixtures and solutions **Avogadro and Dalton** Andrew Norman Meldrum, 1906 **Kinetic Theory**, 19?? *Foundations of Molecular Theory, Comprising Papers and Extracts* John Dalton, Joseph Louis Gay-Lussac, Amedeo Avogadro, 1961 Foundations of the Molecular Theory, 1893 *Molecules and the Molecular Theory of Matter* A. D. Risteen, 2015-06-15 Excerpt from *Molecules and the Molecular Theory of Matter* In the multiplication of popular books on

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