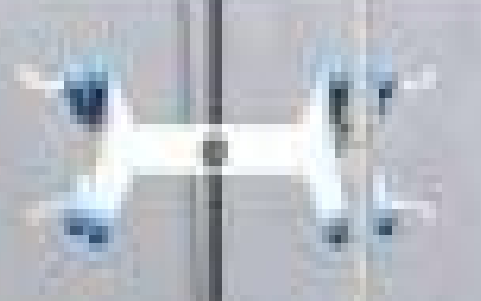


## Containers



Background

Procedures

Lab Notes

Print

By definition, the number of moles of reactant is equal to the volume ( $V$ ) multiplied by the concentration ( $C$ ):

$$n = C \times V$$

From the balanced chemical equation we know that the ratio of the two reacting compounds is 1 to 2:



So at the equivalence point:



Now you can calculate the unknown concentration of the  $\text{H}_2\text{SO}_4$  according to:

$$C_{\text{H}_2\text{SO}_4} = \frac{1}{2} \times \frac{M_{\text{NaOH}} \times V_{\text{NaOH}}}{V_{\text{H}_2\text{SO}_4}}$$

# Late Nite Labs Chemistry

**G Thomas**



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