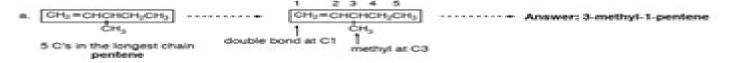
#### Chapter 13 Unsaturated Hydrocarbons

#### Solutions to In-Chapter Problems

- 13.1 To draw a complete structure for each condensed structure, first draw in the multiple bonds. Then draw in all the other C's and H's, as in Example 13.1.
- 13.2 To determine whether each molecular formula corresponds to a saturated hydrocarbon, alkene, or alkyne, recall that the formula for a saturated hydrocarbon is C<sub>0</sub>H<sub>2n-2</sub>, the formula for an alkene is C<sub>0</sub>H<sub>2n-2</sub>.
  - a.  $C_3H_6=C_nH_{2n}=$  alkene c.  $C_5H_{14}=C_nH_{2n-2}=$  alkene d.  $C_6H_{12}=C_nH_{2n-2}=$  alkene
- Use the general formulas [saturated hydrocarbon (C<sub>n</sub>H<sub>2n-2</sub>), alkene (C<sub>n</sub>H<sub>2n</sub>), and alkyne (C<sub>n</sub>H<sub>2n-2</sub>)] to determine the molecular formula for each compound.
  - a. alkene =  $C_nH_{2n}$ ,  $4 \times 2 = 8$ ,  $C_4H_1$ b. saturated hydrocarbon =  $C_nH_{2n-2}$ ,  $(6 \times 2) + 2 = 14$ ,  $C_6H_{14}$ c. alkyne =  $C_nH_{2n-2}$ ,  $(7 \times 2) - 2 = 12$ ,  $C_7H_{12}$ d. alkene =  $C_nH_{2n}$ ,  $5 \times 2 = 10$ ,  $C_7H_{10}$
- 13.4 Give the IUPAC name for each compound using the following steps, as in Example 13.2:
  - [1] Find the longest chain containing both carbon atoms of the multiple bond.
  - [2] Number the chain to give the double bond the lower number.
  - [3] Name and number the substituents and write the complete name.



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#### **Table of Contents Langfield Smith 6e Chapter 13 Solution**

- 1. Understanding the eBook Langfield Smith 6e Chapter 13 Solution
  - The Rise of Digital Reading Langfield Smith 6e Chapter 13 Solution
  - o Advantages of eBooks Over Traditional Books
- 2. Identifying Langfield Smith 6e Chapter 13 Solution
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Langfield Smith 6e Chapter 13 Solution
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Langfield Smith 6e Chapter 13 Solution
  - Personalized Recommendations
  - Langfield Smith 6e Chapter 13 Solution User Reviews and Ratings
  - Langfield Smith 6e Chapter 13 Solution and Bestseller Lists
- 5. Accessing Langfield Smith 6e Chapter 13 Solution Free and Paid eBooks
  - Langfield Smith 6e Chapter 13 Solution Public Domain eBooks
  - Langfield Smith 6e Chapter 13 Solution eBook Subscription Services
  - Langfield Smith 6e Chapter 13 Solution Budget-Friendly Options
- 6. Navigating Langfield Smith 6e Chapter 13 Solution eBook Formats
  - o ePub, PDF, MOBI, and More
  - Langfield Smith 6e Chapter 13 Solution Compatibility with Devices
  - Langfield Smith 6e Chapter 13 Solution Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Langfield Smith 6e Chapter 13 Solution
  - Highlighting and Note-Taking Langfield Smith 6e Chapter 13 Solution
  - Interactive Elements Langfield Smith 6e Chapter 13 Solution
- 8. Staying Engaged with Langfield Smith 6e Chapter 13 Solution

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Langfield Smith 6e Chapter 13 Solution
- 9. Balancing eBooks and Physical Books Langfield Smith 6e Chapter 13 Solution
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Langfield Smith 6e Chapter 13 Solution
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Langfield Smith 6e Chapter 13 Solution
  - Setting Reading Goals Langfield Smith 6e Chapter 13 Solution
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Langfield Smith 6e Chapter 13 Solution
  - Fact-Checking eBook Content of Langfield Smith 6e Chapter 13 Solution
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - o Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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