

Experiment 6B: Calculations with a Chemical Reaction

Objectives:

1. To understand the theory of stoichiometry, including the concept of molar masses and limiting reagents, for homogeneous reaction systems.
2. To calculate the limiting reagent, assuming the starting materials present in the solution.
3. To determine stoichiometrically the maximum yield of a reaction.
4. To calculate the theoretical maximum stoichiometric reagents that could be produced.
5. To compare the theoretical amount of the actual amount of reagents, experimentally calculated yield.

Materials: Please refer to Experiment 6B page 95 of the Practical Chemistry Laboratory

Equipment:

Procedure:

1. Prepare very carefully experiment with the people's instructions.
2. Prepare a solution of 100 ml of water with a concentration of 100 g/l.
3. Calculate 10 ml of concentrated solution of 100 g/l of a solution of 100 g/l.
4. Dilute 100 ml of the solution to 100 ml. Then the concentration of the solution is 100 g/l.
5. Use the 100 ml of the solution to calculate the concentration of the solution.
6. Prepare 100 ml of concentrated solution and use the solution.
7. Use the 100 ml of the solution to calculate the concentration of the solution. Then the concentration is 100 g/l.
8. Use the concentration of the solution to calculate the concentration of the solution.
9. Prepare the solution of a solution of 100 g/l of a solution of 100 g/l.
10. Calculate the concentration of the solution of the solution of 100 g/l of a solution of 100 g/l.

Lab 6b Calculations With A Chemical Reaction Answers

JE Gale



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