

Part I: Determining the calorimeter constant:

Data	Trial	Trial
(a) Mass of empty Styrofoam cups	8.6 g	8.6 g
(b) Mass of cups + 70 mL water	76.6 g	77.7 g
(c) Mass of cups + 70 mL water + 30 mL hot	103.4 g	104.0 g
(d) Initial temperature of water in the calorimeter	18.5 °C	21.0 °C
(e) Temperature of the boiling water bath	98.1 °C	98.1 °C
(f) Final temperature of calorimeter + added	39.8 °C	40.6 °C

Calculation		
(g) Mass of cool water in cups, $m_{CW} [b - a]$	68.0 g	g
(h) Mass of added hot water, $m_{HW} [c - b]$	26.8 g	g
(i) Temperature change of cool water in the calorimeter ($\Delta T_{CW} = T_f - T_{i(CW)}$)	+ 21.3 K	+ K
(j) Temperature change of added hot water ($\Delta T_{HW} = T_f - T_{i(HW)}$)	- 58.3 K	- K
(k) Calculate the heat <i>lost</i> by the hot water $\Delta H_{HW} = m_{HW} C_s \Delta T_{HW}$	- 6530 J	- J
(l) Calculate the heat <i>gained</i> by the cold water $\Delta H_{CW} = m_{CW} C_s \Delta T_{CW}$	+ 6050 J	+ J
(m) Calculate the heat <i>gained</i> by the calorimeter $\Delta H_{cal} = -\Delta H_{HW} - \Delta H_{CW}$	+ 480 J	+ J
(n) Calculate the calorimeter constant, B $B = \Delta H_{cal} / \Delta T_{CW}$	22.5 J/K	J/K
(o) Average calorimeter constant, B		J/K

Lab 22 Specific Heat Of Metals Answers

Padhraic Smyth

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