

2015

Subject : Chemistry (Honours)

Paper : V

Time: 2 Hours

Full Marks: 50

*The figures in the margin indicate full marks.
Candidates are required to give their answers in their own words
as far as practicable.*

1. Answer any four questions:

2×4=8

- Depict the ambidentate and flexidentate behaviours of thiocyanate anion.
- SiCl_4 has lower boiling point than CCl_4 . Explain.
- Cite and draw one example each of $3c-2e$ and $2c-2e$ bonds.
- Briefly comment on the compound K_3C_{60} .
- Write down the structural formulae of sodium pyrophosphate and sodium metaphosphate.
- What is meant by chalcogenide? Give example.
- 'NaOH is base but ClOH is acid' - Give reason.

2. Answer any three questions:

6×3=18

- What are inner metallic complexes? Give example. 2
 - Chromium(III) chloride forms three different hydrates of the same mole ratio $\text{Cr:Cl:H}_2\text{O}=1:3:6$. The violet form does not lose water over conc. H_2SO_4 and gives 3 equivalent of AgCl on treatment with AgNO_3 . The other two forms are green and lose 1 and 2 moles of H_2O over conc. H_2SO_4 and give 2 and 1 equivalent of AgCl , respectively with AgNO_3 . Write down the structures of the three isomeric coordination compounds and explain the isomerism involved. 4
- Define and exemplify chelate complex. 3
 - 'The stability of $[\text{Ni}(\text{en})_3]^{2+}$ (en =ethylenediamine) is much greater than that of $[\text{Ni}(\text{NH}_3)_6]^{2+}$, although both contain six Ni-N bonds.' - Explain. 3
- Apply radius ratio principle to explain the difference between crystal structures of NaCl and CsCl . 2
 - ' MgO is a stable ionic solid though large amount of energy is required to form Mg to Mg^{2+} and to dissociate O_2 into atomic oxygen and thereafter to form O^{2-} ion.' - Account for the fact. 2

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Table of Contents General Chemistry Question 2014 1st Year Burdwan University

1. Understanding the eBook General Chemistry Question 2014 1st Year Burdwan University
 - The Rise of Digital Reading General Chemistry Question 2014 1st Year Burdwan University
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Question 2014 1st Year Burdwan University
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Question 2014 1st Year Burdwan University
 - User-Friendly Interface

4. Exploring eBook Recommendations from General Chemistry Question 2014 1st Year Burdwan University
 - Personalized Recommendations
 - General Chemistry Question 2014 1st Year Burdwan University User Reviews and Ratings
 - General Chemistry Question 2014 1st Year Burdwan University and Bestseller Lists
5. Accessing General Chemistry Question 2014 1st Year Burdwan University Free and Paid eBooks
 - General Chemistry Question 2014 1st Year Burdwan University Public Domain eBooks
 - General Chemistry Question 2014 1st Year Burdwan University eBook Subscription Services
 - General Chemistry Question 2014 1st Year Burdwan University Budget-Friendly Options
6. Navigating General Chemistry Question 2014 1st Year Burdwan University eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Question 2014 1st Year Burdwan University Compatibility with Devices
 - General Chemistry Question 2014 1st Year Burdwan University Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Question 2014 1st Year Burdwan University
 - Highlighting and Note-Taking General Chemistry Question 2014 1st Year Burdwan University
 - Interactive Elements General Chemistry Question 2014 1st Year Burdwan University
8. Staying Engaged with General Chemistry Question 2014 1st Year Burdwan University
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Question 2014 1st Year Burdwan University
9. Balancing eBooks and Physical Books General Chemistry Question 2014 1st Year Burdwan University
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Question 2014 1st Year Burdwan University
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Question 2014 1st Year Burdwan University
 - Setting Reading Goals General Chemistry Question 2014 1st Year Burdwan University
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of General Chemistry Question 2014 1st Year Burdwan University
 - Fact-Checking eBook Content of General Chemistry Question 2014 1st Year Burdwan University
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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