

Unit 5: Chemical Kinetics Exam

1.

Trial	$[A_2]$	$[B]$	Initial rate ($M s^{-1}$)
1	0.10	0.50	2.5×10^{-3}
2	0.20	0.50	5.0×10^{-3}
3	0.30	0.05	5.0×10^{-3}
4	0.30	0.10	1.0×10^{-2}

An experiment was conducted to determine the rate law for the reaction $A_2(g) + B(g) \rightarrow A_2B(g)$. The table above shows the data collected. Based on the data in the table, which statement is correct?

- (A) Since the rate law can be expressed as **rate** = $k[A_2]^2$, tripling the concentration of A_2 will cause a 9-fold increase in the rate of the reaction.
- (B) Since the rate law can be expressed as **rate** = $k[A_2][B]$, doubling the concentrations of A_2 and B will quadruple the rate of the reaction.
- (C) Since the rate law can be expressed as **rate** = $k[A_2]^2[B]$, tripling the concentration of A_2 while keeping the concentration of B constant will triple the rate of the reaction.
- (D) Since the rate law can be expressed as **rate** = $k[A_2][B]^2$, doubling the concentration of B while keeping the concentration of A_2 constant will double the rate of the reaction.

2. $2 NO(g) + Cl_2(g) \rightarrow 2 NOCl(g)$

Experiment	$[NO]$	$[Cl_2]$	Initial Rate of $NOCl$ Formation ($M s^{-1}$)
1	0.0250	0.0250	4.55×10^{-3}
2	0.0250	0.1000	9.10×10^{-3}
3	0.0500	0.0250	1.82×10^{-2}

The initial rates of the reaction represented by the equation shown above were measured for different initial concentrations of $NO(g)$ and $Cl_2(g)$. Based on the data given in the table above, which of the following is the rate law expression for the reaction, and why?

- (A) **Rate** = $k[NO]^2$, because the initial rate quadrupled when $[NO]$ was doubled but remained constant when $[Cl_2]$ was doubled.
- (B) **Rate** = $k[NO][Cl_2]$, because the initial rate doubled when either $[NO]$ or $[Cl_2]$ was doubled.
- (C) **Rate** = $k[NO][Cl_2]^2$, because the initial rate doubled when $[NO]$ was doubled and quadrupled when $[Cl_2]$ was doubled.
- (D) **Rate** = $k[NO]^2[Cl_2]$, because the initial rate quadrupled when $[NO]$ was doubled and doubled when $[Cl_2]$ was doubled.

General Chemistry Ii Chapter Chemical Kinetics Exam

Zhe Xu



General Chemistry II Chapter Chemical Kinetics Exam:

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