

General Chemistry 10th

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CHAPTER 1

Chemistry and Measurement

■ SOLUTIONS TO EXERCISES

Note on significant figures: If the final answer to a solution needs to be rounded off, it is given first with one nonsignificant figure, and the last significant figure is underlined. The final answer is then rounded to the correct number of significant figures. In multistep problems, intermediate answers are given with at least one nonsignificant figure; however, only the final answer has been rounded off.

- 1.1. From the law of conservation of mass,

$$\text{Mass of wood} + \text{mass of air} = \text{mass of ash} + \text{mass of gases}$$

Substituting, you obtain

$$1.85 \text{ grams} + 9.45 \text{ grams} = 0.28 \text{ grams} + \text{mass of gases}$$

or,

$$\text{Mass of gases} = (1.85 + 9.45 - 0.28) \text{ grams} = 11.02 \text{ grams}$$

Thus, the mass of gases in the vessel at the end of the experiment is 11.02 grams.

- 1.2. Physical properties: soft, silvery-colored metal; melts at 64°C.
Chemical properties: reacts vigorously with water, with oxygen, and with chlorine.

- 1.3. a. The factor 9.1 has the fewest significant figures, so the answer should be reported to two significant figures.

$$\frac{5.61 \times 7.891}{9.1} = 4.\underline{86} = 4.9$$

- b. The number with the least number of decimal places is 8.91. Therefore, round the answer to two decimal places.

$$8.91 - 6.435 = 2.\underline{475} = 2.48$$

- c. The number with the least number of decimal places is 6.81. Therefore, round the answer to two decimal places.

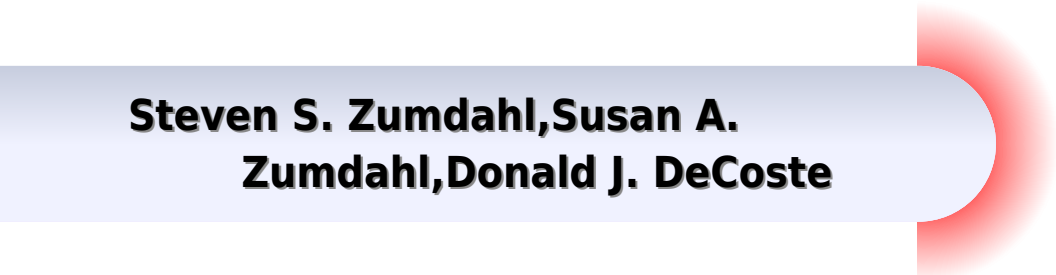
$$6.81 - 6.730 = 0.\underline{080} = 0.08$$

- d. You first do the subtraction within parentheses. In this step, the number with the least number of decimal places is 6.81, so the result of the subtraction has two decimal places. The least significant figure for this step is underlined.

$$38.91 \times (6.81 - 6.730) = 38.91 \times 0.\underline{080}$$

General Chemistry Ebbing 10th Edition Solutions

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Zumdahl, Donald J. DeCoste**



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Student Solutions Manual for Ebbing/Gammon's General Chemistry, 11th Darrell Ebbing, Steven D. Gammon, 2016-02-09
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