

1. According to the Brønsted—Lowry definition, which chemical species can function *both* as an acid and as a base?

- Cl^-
- SO_4^{2-}
- NH_4^+
- HCO_3^-
- H_3O^+

2. In the reaction $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$ which is an acid—base conjugate pair?

- H_2O and HCN
- H_2O and OH^-
- CN^- and H_2O
- HCN and OH^-

3. Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX

- It is less basic than a 1 M solution of NaY.
- It is more basic than a 1 M solution of NaY.
- It yields a neutral solution.
- It is more concentrated than a 1 M solution of NaY.

4. The pH of a 0.03 M HCl solution is

a.	1.5
b.	2.5
c.	3.5
d.	12.5

5. The pH of a 1.0×10^{-3} M $\text{Ba}(\text{OH})_2$ solution at 25°C is

a.	2.7
b.	3.0
c.	11.0
d.	11.3

4. Which is the strongest acid?

- HClO
- HClO_2
- HClO_3
- HClO_4

7. Which salt reacts with water (hydrolyzes) to produce a basic solution?

a. $\text{NaC}_2\text{H}_3\text{O}_2$
b. NaNO_3
c. NH_4Cl
d. BaSO_4

8. The weakest of the bases listed is

Acid	Conjugate Base	K _a (Ionization Constant of Acid)
HCl	Cl ⁻	100% ionized
HSO ₄ ⁻	SO ₄ ²⁻	1.2 x 10 ⁻²
H ₂ S	HS ⁻	5.7 x 10 ⁻⁸
HS ⁻	S ²⁻	1.2 x 10 ⁻¹⁴

a.	Cl^-
b.	CN^-
c.	HS^-
d.	S^{2-}
e.	SO_4^{2-}

Number of hauls	<i>M. undulatus</i> (%)	<i>M. setulosus</i> (%)	<i>M. setulosus</i> + <i>M. undulatus</i> (%)
1	10	10	20
2	30	20	50
3	50	30	80
4	70	40	100
5	85	50	100
6	95	55	100
7	100	60	100
8	100	65	100
9	100	70	100
10	100	75	100

10. Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate $\text{KC}_2\text{H}_3\text{O}_2$ is less alkaline than a 0.10 M solution of potassium cyanide, KCN?

- Hydrocyanic acid is a weaker acid than acetic acid.
- Hydrocyanic acid is less soluble in water than acetic acid.
- Cyanides are less soluble than acetates.
- Acetic acid is a weaker acid than hydrocyanic acid.

11. In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of

- 0.100 M sodium solution,
- 0.0500 M sodium hydroxide solution,
- 0.0500 M benzoic acid solution,
- 0.0500 M sodium benzoate solution.

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