

## CHEM 102 – Recitation Week 9

Using the “*Key Concepts*” and the “*Key Equations and Relationships*” in the *Chapter Summary* at the end of each chapter, identify and type in this Word document some concepts and/or equations needed to solve each item. Be prepared to share your ideas in class discussions. You are NOT required to solve the problems in detail. If you choose to do it to better participate in class discussion, keep in mind that full problem-solving is not graded for the pre-recitation assignment. To get full recitation credit, you must **identify at least 1-2 correct and also different concepts for each question** (see the *Answer Key* posted for week 1, which shows comprehensive examples of concept identification). **Copying and pasting very general concepts (such as, “units of measure” or “dimensional analysis”, etc.) as the only identified concepts for multiple questions is not allowed.**

**Ch. 17, 92.** The solubility of copper(I) chloride,  $\text{CuCl}$ , is 3.91 mg per 100.0 mL of solution. Calculate  $K_{sp}$  for  $\text{CuCl}$ .

**Concepts needed:** Using the formula for  $K_{sp} = [\text{Cu}^+][\text{Cl}^-]$ , solve for  $K_{sp}$ . Understanding how to calculate  $K_{sp}$  when given mg per mL of solution.

**Ch. 17, 93.** Calculate the molar solubility of barium fluoride,  $\text{BaF}_2$ , in each liquid or solution below. For barium fluoride,  $K_{sp} = 2.45 \times 10^{-6}$ .

- a) pure water
- b) 0.10 M  $\text{Ba}(\text{NO}_3)_2$

**Concepts needed:** Molar solubility is essentially the solubility of a solution in moles/liter. To calculate molar solubility, you have to set up an ICE table and use  $K_{sp}$  to find each concentration.

**Ch. 18, 36.** Without doing any calculations, determine the sign of  $\Delta G^\circ$  for each chemical reaction.

- a)  $\text{Mg}(s) + \text{Cl}_2(g) \rightarrow \text{MgCl}_2(s)$
- b)  $2\text{H}_2\text{S}(g) + 3\text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(g) + 2\text{SO}_2(g)$

# Futa Chem 102

**Raffaela Di Napoli**



**Futa Chem 102:**

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