Gas Laws Worksheet

atm = 760.0 mm Hg = 101.3 kPa= 760.0 torr

Boyle's Law Problems:

 If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(748-4)(22.54) = (725-4) V_2$$

 $V_2 = \frac{(748-4)(22.54)}{(725-4)}$

2 A gas with a volume of 4.0L at a pressure of 205kPa is allowed to expand to a volume of 12.0L. What is the pressure in the container if the temperature remains constant?

 What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

4 A 40 0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

Exploring Gas Laws Worksheet Answers

JS Bruner

Exploring Gas Laws Worksheet Answers:

Whispering the Techniques of Language: An Psychological Journey through Exploring Gas Laws Worksheet Answers

In a digitally-driven earth wherever displays reign great and immediate interaction drowns out the subtleties of language, the profound secrets and mental subtleties concealed within phrases frequently get unheard. Yet, nestled within the pages of **Exploring Gas Laws Worksheet Answers** a captivating fictional treasure pulsating with raw thoughts, lies an extraordinary journey waiting to be undertaken. Composed by a talented wordsmith, that charming opus invites visitors on an introspective journey, lightly unraveling the veiled truths and profound impact resonating within ab muscles material of every word. Within the psychological depths of this moving evaluation, we shall embark upon a genuine exploration of the book is primary themes, dissect their interesting writing fashion, and fail to the effective resonance it evokes deep within the recesses of readers hearts.

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