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THE NATURE OF GASES

Section Review

Objectives

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory.
- Define the relationship between Kelvin temperature and average kinetic energy

Vocabulary

- kinetic energy
- kinetic theory:
- gas pressure.

- Vacuum
- atmospherie pressure
- barometer

- pascal (Pa):
- · standard atmosphere (atm)

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

				se particles of gas are relatively
				between particles are perfectly
				the simultaneous collisions of billions of
	particles(with an object.	Barometers are used to measur	repress@re.	Standard conditions are defined as a
	temperature of	and a pressure of		
ı.			8	
2.			9	
3.	10	<u> </u>		
4.				
10				
10.				

- ____ 11. Atmospheric pressure is 760 mm Hg.
- _____ 12. The SI unit of pressure is the pascal.

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