

CHAPTER 8

CHEMICAL EQUILIBRIUM MCQs

- Q.1 A reaction is reversible because
(a) reactants are reactive (b) products are reactive
(c) products are stable (d) reactants are stable
- Q.2 A large value of K_c means that at equilibrium
(a) less reactants and more products
(b) more reactants and less product
(c) same amount
(d) none
- Q.3 Extent to $H_2 + I_2 \rightarrow 2HI$ can be increased by
(a) increasing pressure (b) increasing product
(c) increasing temp (d) adding a catalyst
- Q.4 Strength of an acid can be determined by
(a) PK_a (b) PK_p
(c) POH (d) PK_w
- Q.5 In an exothermic reversible reaction increase in temp shifts the equilibrium to
(a) reactant side (b) product side
(c) remains unchanged (d) none
- Q.6 Units of K_w are
(a) mole dm^{-3} (b) $\text{mole}^2 \text{ dm}^{-3}$
(c) $\text{mole}^2 \text{ dm}^{-6}$ (d) $\text{mole}^2 \text{ dm}^{-3}$
- Q.7 A basic Buffer solution can be prepared by mixing
(a) weak acid and its salt with strong base
(b) strong acid and its salt with weak base
(c) weak base and its salt with strong acid
(d) strong base and its salt with weak acid
- Q.8 Buffer action can be explained by
(a) common ion effect (b) law of mass action

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