

Laboratory Exercise: Double Replacement Reactions...

TABLE II: SCALE UP OF PRECIPITATION REACTIONS

Row #	Chemical Formula of Reagent	Chemical Formula of Reagent	List the four ions present in reagents		insoluble	soluble
					Formula of Solid Product appearance in test tube	Formula of Aqueous Product appearance in test tube
2	NaOH	FeCl ₃	Na ⁺	OH ⁻	Fe(OH) ₃	
			Fe ³⁺	Cl ⁻	very small orange	very cloudy orange
3	NaNO ₃	FeCl ₃	Na ⁺	NO ₃ ⁻		
			Fe ³⁺	Cl ⁻	solid white powder	yellow cloudy
5	Na ₂ CO ₃	FeCl ₃	Na ⁺	CO ₃ ²⁻		
			Fe ³⁺	Cl ⁻	solid orange powdery	very cloudy orange
6	Na ₃ PO ₄	FeCl ₃	Na ⁺	PO ₄ ³⁻		
			Fe ³⁺	Cl ⁻	yellowish small	yellow cloudy
12	AgNO ₃	NaOH				
					not almost brown	cloudy brown
13	NaOH	CuSO ₄	N			
					gel like chunk	blue uncent
17	AgNO ₃	Na ₂ CO ₃				
					white flake small	white cloudy
18	AgNO ₃	Na ₃ PO ₄				
					yellow egg	clear yellow
19	CuSO ₄	Na ₂ CO ₃				
					small blue/white powder	cloudy blue
20	CuSO ₄	Na ₃ PO ₄				
					very small blue particles	cloudy light-blue

Lab 9 2 Double Replacement Reactions

B Lingard



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