

### Biomorphic Formulation Practical assignment

1. Determine molecular formula of the compound. The following is report of analysis:

$$\begin{aligned} \text{C} &= 69.57\% & \text{H} &= 4.91\% & \text{N} &= 14.52\% \end{aligned}$$

A sample of compound is subjected to elemental determination in a mixture. The percentage obtained for nitrogen is 14.52% and the atomic weight of nitrogen is 14.01. Determine the molecular formula of the compound.

Mass of carbon in 100 g	69.57 g
Mass of hydrogen in 100 g	4.91 g
Mass of nitrogen in 100 g	14.52 g

2. Calculate the molar mass of the compound from the data given in the table.

$$\begin{aligned} \text{Molar mass of } \text{C}_6\text{H}_5\text{N} &= 6 \times 12 + 5 \times 1 + 14 \\ &= 77 \text{ g/mol} \end{aligned}$$

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4. Write the structural formula of the compound.



5. Calculate the empirical formula of the compound.



# Lab 6 1 Empirical Formula Answers

**Mr. Rohit Manglik**



## **Lab 6 1 Empirical Formula Answers:**

## Reviewing **Lab 6 1 Empirical Formula Answers**: Unlocking the Spellbinding Force of Linguistics

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