

Ionic Salts Simulation Lab**Learning Goals:**

- Understand how an ionic salt dissolves describing a saturated solution microscopically and macroscopically.
- Calculate solubility in grams/100ml
- Determine why some salts are more soluble than others
- Design and carry out a virtual lab to test a hypothesis

Directions: Go to

http://phet.colorado.edu/simulations/sims.php?sim=Salts_and_Solubility and run the Salts and Solubility simulation. Take a few minutes to get use to how the simulation works. Then answer the questions below.

Questions:

- 1) Describe what you observe when you shake table salt into water. Test numbers of particles with different volumes.
 - a. When the salt dissolves, which element becomes the anion (the negatively charged atom) and which one becomes the cation (the positively charged atom)?
 - b. What is the relationship between amount of particles and amount of water? (direct or indirect).
 - c. At 5.00×10^{-23} what is the point of saturation (point where you can't dissolve any more salt)?
- 2) Fill in the chart below to see if the ratio of particles to water is constant at the saturation point.

Particles	Water Volume	Ratio (particles/volume)

- a. Calculate the constant below (show your work)

- 3) Test a variety of slightly soluble salts, noting the relationship between numbers of each ion pair. Fill in the chart below to help you see a relationship.
 - a. How many particles of each element need to be present before the salt stops dissolving (point of saturation)?

Lab 30 Solubility Of A Salt Answers

Benjamin Lillard



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Journal of the Society of Chemical Industry , 1891 *CliffsNotes AP Chemistry* Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book **Examination Papers Set at the Certificate Examination for Teachers in Elementary Schools** Great Britain. Board of Education, 1911

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