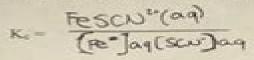
Report Sheet - Experiment 13-Chemical Equilibrium

- 1. Write the chemical equation for the reversible reaction studied in this experiment:

 Fe *(aq)+ SCN * > Fe (SCN) *(aq)
- 2. Write the equilibrium expression for the reaction studied in this experiment:



3. Complete the following table from your observations of test tubes #1-#6:

Test Tube	Change	Initial Color	Final Color (darker/lighter)
	Initial color before anything is added		
1	Added Fe(NO ₃) ₃	Reddish color	dorker
2	Added KSCN	Reddish color	darker
3	Added NH ₄ CI	Reddish color	Lighter
4	Added AgNO	Reddish color	Lighter
5	Cold	Reddish color	darker
6	Hot	Reddish color	lighter

4. Complete the following table for test tubes #2 - 4:

Test Tube	What concentration was changed and how it was changed?	The equilibrium shifted to left/right/no change	What evidence for the direction of shift do you have?
2	Fesch t is a mix of an iron I relaish color + thiconate has no color	Shifted Paykt	principle but react on from lett to right accus was when more reactant add
3	(SCN) was	Shifted Right	the product goes up
4	Fezt concentration was	Shifetell Just+	Fe COL WILL Form which leads to a lower Fe to concentration

Lab 1rate And Equilibrium Report Answer

IM Harris

Lab 1rate And Equilibrium Report Answer:

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