

1. Write the chemical equation for the reversible reaction studied in this experiment:



2. Write the equilibrium expression for the reaction studied in this experiment:

$$K_c = \frac{\text{FeSCN}^{2+}(\text{aq})}{[\text{Fe}^{2+}]_{\text{aq}}[\text{SCN}^{-}]_{\text{aq}}}$$

3. Complete the following table from your observations of test tubes #1-#6:

Test Tube	Change	Initial Color	Final Color (darker/lighter)
	Initial color before anything is added		
1	Added $\text{Fe}(\text{NO}_3)_3$	Reddish color	darker
2	Added KSCN	Reddish color	darker
3	Added NH_4Cl	Reddish color	lighter
4	Added AgNO_3	Reddish color	lighter
5	Cold	Reddish color	darker
6	Hot	Reddish color	lighter

4. Complete the following table for test tubes #2 - 4:

Test Tube	What concentration was changed and how it was changed?	The equilibrium shifted to left/right/no change	What evidence for the direction of shift do you have?
2	FeSCN^{2+} is a mix of an iron/reddish color + thiocyanate has no color	Shifted Right	The Le Chatelier principle w/ reaction from left to right occurs when more reactant was added.
3	$[\text{SCN}^{-}]$ was increased	Shifted Right	As more reactant was added, the concentration of the product goes up
4	Fe^{2+} concentration was changed	Shifted Left	FeCO_3 will form which leads to a lower Fe^{2+} concentration

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LP Steffe



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