TYPES OF CHEMICAL REACTIONS LAB PART#1

 Purpose: To view the actual chemical reactions, write the correct balanced chemical equation, and type of chemical reaction.

Connection to Business: Businesses utilize many types of chemical reactions to produce, maintain, and sterilize their product or to prevent pollution.

II. Materials: burner, test tube, metal tongs (only...no rubber ends), microspatula, evaporating dish, wood splints

Chemicals: copper (wire), magnesium (ribbon), copper (II) carbonate SAFETY GOGGLES!

- III. Procedure: Part #1
 - 1. Obtain materials.
 - 2. Use the fine sandpaper to clean a piece of copper wire until shiny.
 - 3. Set up a Bunsen burner.
 - Using tongs, hold the wire in the hottest part of the burner flame for 1-2 minutes. Record your observation of the wire's appearance afterwards.
 - Place an evaporating dish near the Bunsen burner and repeat step #4 except place the remaining magnesium into the evaporating dish. (WARNING: AVOID LOOKING DIRECTLY AT THE FLAME)
 - Place 2 heaping microspatulas of copper (II) carbonate into a clean, dry test tube.
 - Using a test tube holder, hold the test tube (at a 45° angle pointed away from people) over the Bunsen burner flame for 3 minutes or until completely black.
 - (Allow this to cool slightly before placing into the wood test tube holder.)

 8. Light a wood splint and place inside the test tube. Record your observations.
 - Dispose of all the metals and burnt powders in a zip-lock bag and place in the trash.
 - 10. Clean and return all lab equipment.
- IV. Data/Questions: Part #1 Data Table

Sample	Appearance Before Reaction	Appearance After Reaction
1. Copper	Answer: shiny, reddish brown	Answer: black, powder coating
2. Magnesium	Answer: silver, tinsel	Answer: black ashes, bright light
3. Mercury (II) Oxide (will not perform experiment)	Reddish powder	Liquid mercury formed; flame burned brighter when placed in the test tube
4. Copper (II) Carbonate	Answer: green powder	Answer: decomposes into gas CO ₂ and black residue

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