Review for the Full Year General Chemistry American Chemical Society Exam (a) Thermodynamics, (b) Transition Metals, (c) Redox and Electrochemistry, (d) Nuclear and (e) Organic

THERMODYNAMICS

- 84. A 10.0 g sample of silver is heated to 100.0 "C and then added to 20.0 g of water at 23.0 "C in an insulated calorimeter. At thermal equilibrium the temperature of the system was measured as 25.0 °C. What is the specific heat of silver? [specific heat = 4.2 J g-1 "C-1]
 - (A) 0.11 J/g K (C) 17 J/g K
- (B) 0.22 J/g K (D) 34 J/g K
- 85. What is AH in kJ for the reaction:

[ΔH_I^D (kJ/mol):C₀H_{2/gl} +227; H₂O_{(gl} -242; CO-min -3930

- (A) -1830 kJ
- (B) -2284 kJ
- (D) -1605 kJ (C) -2510 kJ
- 86. Given these equations

calculate the heat of formation of SO_{Mill}.

- (A) -500 kJ mol-1
- (C) +100 kJ mol⁻¹
- (8) -400 kJ mol-1
- (ID) +2:00 kJ mol-1

87. Given:

How much energy is produced in the combustion of 28 g of CO? (Atomic Masses: C 12.0 g mol* 1; O 16.0 g mol-1]

- (A) 88 kJ
- (B) 109 kJ
- (C) 175 kJ (D) 284 kJ
- 88. Calculate the value of AH (in kJ mol 1) for the reaction.

Bond Energies (kJ molt⁻¹) H-H 435 NeN 946 (in No) N-H 389

- (A) 2340 kJ of heat absorbed
- (B) 213 kJ of heat absorbed
- (C) 2340 kJ of heat evolved
- (D) 83 kJ of heat evolved

- 89. Which change is likely to be accompanied by the greatest increase in entropy?

- (C) CO₂(s) ----- CO₂(g) (at -70 °C)
- (D) H2O(8) ------ H2O(8 (at 100 °C)
- 90. For which process is the entropy change per mole the largest at constant temperature?
 - (A) H₂O₁₀ ---- H₂O₁₀₀
 - (B) H₂O_(N) ----- H₂O_(S)
 - (C) H₂O_(a) → H₂O_(b)
 - (D) H₂O₍₀₎ → H₂O₍₁₀₎
- 91. In which process is entropy decreased?
 - (A) dissolving sugar in water
 - (B) expanding a gas
 - (C) evaporating a liquid
 - (D) freezing water
- 92. When Al₂O_{3/32} is formed from the elements at standard conditions, the values of ΔH^0 and ΔG^0 at 298 K are
 - -1676 kJ mol-1 and -1577 kJ mol-1. respectively. The standard entropy of formation per mole, in joules per degree, will be
 - (A) 332(D) -0.0933
- (B) -157 (05) + 15.7
- (C) -93.3
- 93. Vaporization of a liquid is an example of a process for which
 - (A) AH, AS, and AG are positive at all temperatures.
 - (B) \(\Delta H \) and \(\Delta S \) are positive.
 - (C) AG is negative at low temperatures. positive at high temperatures.
 - $(D) \Delta H = \Delta S$
- 94. A particular chemical reaction has a negative AH and negative AS. Which statement is correct?
 - (A) The reaction is spontaneous at all termperatures.
 - (B) The reaction is nonspontaneous at all beimperatures.
 - (C) The reaction becomes spontaneous as temperature increases.
 - (D) The reaction becomes spontaneous as temperature decreases.

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