

**Review for the Full Year General Chemistry  
American Chemical Society Exam**

**(a) Thermodynamics, (b) Transition Metals, (c) Redox and Electrochemistry, (d) Nuclear and (e) Organic**

**THERMODYNAMICS**

84. A 10.0 g sample of silver is heated to 100.0 °C and then added to 20.0 g of water at 23.0 °C in an insulated calorimeter. At thermal equilibrium the temperature of the system was measured as 25.0 °C. What is the specific heat of silver? [specific heat =  $4.2 \text{ J g}^{-1} \text{ }^{\circ}\text{C}^{-1}$ ]

- (A) 0.11 J/g K (C) 17 J/g K  
(B) 0.22 J/g K (D) 34 J/g K

85. What is  $\Delta H$  in kJ for the reaction:



- (A) -1830 kJ (B) -2284 kJ  
(C) -2510 kJ (D) -1605 kJ

86. Given these equations:



calculate the heat of formation of  $\text{SO}_3(\text{g})$ .

- (A) -500 kJ mol<sup>-1</sup> (C) +100 kJ mol<sup>-1</sup>  
(B) -400 kJ mol<sup>-1</sup> (D) +200 kJ mol<sup>-1</sup>

87. Given:



How much energy is produced in the combustion of 28 g of CO? [Atomic Masses: C 12.0 g mol<sup>-1</sup>; O 16.0 g mol<sup>-1</sup>]

- (A) 88 kJ (B) 109 kJ  
(C) 175 kJ (D) 284 kJ

88. Calculate the value of  $\Delta H$  (in kJ mol<sup>-1</sup>) for the reaction:



Bond Energies (kJ mol <sup>-1</sup> )	
H-H	435
N=N	945 (in N <sub>2</sub> )
N-H	389

- (A) 2340 kJ of heat absorbed  
(B) 213 kJ of heat absorbed  
(C) 2340 kJ of heat evolved  
(D) 83 kJ of heat evolved

89. Which change is likely to be accompanied by the greatest increase in entropy?



90. For which process is the entropy change per mole the largest at constant temperature?



91. In which process is entropy decreased?

- (A) dissolving sugar in water  
(B) expanding a gas  
(C) evaporating a liquid  
(D) freezing water

92. When  $\text{Al}_2\text{O}_3(\text{s})$  is formed from the elements at standard conditions, the values of  $\Delta H^{\circ}$  and  $\Delta G^{\circ}$  at 298 K are -1676 kJ mol<sup>-1</sup> and -1577 kJ mol<sup>-1</sup>, respectively. The standard entropy of formation per mole, in joules per degree, will be

- (A) -332 (B) -157 (C) -93.3  
(D) -0.0933 (E) +15.7

93. Vaporization of a liquid is an example of a process for which

- (A)  $\Delta H$ ,  $\Delta S$ , and  $\Delta G$  are positive at all temperatures.  
(B)  $\Delta H$  and  $\Delta S$  are positive.  
(C)  $\Delta G$  is negative at low temperatures, positive at high temperatures.  
(D)  $\Delta H = \Delta S$

94. A particular chemical reaction has a negative  $\Delta H$  and negative  $\Delta S$ . Which statement is correct?

- (A) The reaction is spontaneous at all temperatures.  
(B) The reaction is nonspontaneous at all temperatures.  
(C) The reaction becomes spontaneous as temperature increases.  
(D) The reaction becomes spontaneous as temperature decreases.

# General Chemistry Exam 4

**LP Steffe**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'LP Steffe' text.

## **General Chemistry Exam 4:**

### **Peterson's Grad Programs in Physical Sciences, Math, Ag Sciences, Envir & Natural Res 20154 (Grad 4)**

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population are equally beneficial for all students While stressing that there are many ways to teach effectively and that readers can be flexible in picking and choosing among the strategies she presents Sandra McGuire offers the reader a step by step process for delivering the key messages of the book to students in as little as 50 minutes Free online supplements provide three slide sets and a sample video lecture This book is written primarily for faculty but will be equally useful for TAs tutors and learning center professionals For readers with no background in education or cognitive psychology the book avoids jargon and esoteric theory

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**The Chemical News and Journal of Industrial Science**, 1912      *Survival Guide to Organic Chemistry* Patrick E. McMahon, Bohdan B. Khomtchouk, Claes Wahlestedt, 2016-12-19 Reviews key general chemistry concepts and techniques adapted for application to important organic principles Provides practical guidance to help students make the notoriously well known and arduous transition from general chemistry to organic chemistry Explains organic concepts and reaction mechanisms generally expanding the focus on how to understand each step from a more intuitive viewpoint Covers concepts that need further explanation as well as those that summarize and emphasize key ideas or skills necessary in this field An added bonus is help with organizing principles to make sense of a wide range of similar reactions and mechanisms Implements a user friendly process to achieve the end result of problem solving Covers organic chemistry I and II concepts at the level and depth of a standard ACS organic chemistry curriculum features practice problems and solutions to help master the material including an extensive and comprehensive bank of practice exams with solutions      **Peterson's Graduate**

**Programs in the Physical Sciences 2011** Peterson's, 2011-05-01 Peterson's Graduate Programs in the Physical Sciences contains a wealth of information on colleges and universities that offer graduate work in Astronomy and Astrophysics Chemistry Geosciences Marine Sciences and Oceanography Meteorology and Atmospheric Sciences and Physics The institutions listed include those in the United States Canada and abroad that are accredited by U S accrediting bodies Up to date information collected through Peterson's Annual Survey of Graduate and Professional Institutions provides valuable information on degree offerings professional accreditation jointly offered degrees part time and evening weekend programs postbaccalaureate distance degrees faculty students degree requirements entrance requirements expenses financial support faculty research and unit head and application contact information As an added bonus readers will find a helpful See Close Up link to in depth program descriptions written by some of these institutions These Close Ups offer detailed information about the physical sciences program faculty members and their research and links to the program or department's Web site In addition there are valuable articles on financial assistance and support at the graduate level and the graduate admissions process with special advice for international and minority students Another article discusses important facts about

accreditation and provides a current list of accrediting agencies

**The Handbook of Research Synthesis** Harris Cooper, Larry V. Hedges, 1993-11-23 The Handbook is a comprehensive treatment of literature synthesis and provides practical advice for anyone deep in the throes of just teetering on the brink of or attempting to decipher a meta analysis Given the expanding application and importance of literature synthesis understanding both its strengths and weaknesses is essential for its practitioners and consumers This volume is a good beginning for those who wish to gain that understanding

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