

Unit 5: Chemical Kinetics Exam

1.

Trial	$[A_2]$	$[B]$	Initial rate (M/s)
1	0.10	0.50	2.5×10^{-3}
2	0.20	0.50	5.0×10^{-3}
3	0.30	0.25	5.0×10^{-3}
4	0.30	0.10	1.0×10^{-3}

An experiment was conducted to determine the rate law for the reaction $A_2(g) + B(g) \rightarrow A_2B(g)$. The table above shows the data collected. Based on the data in the table, which statement is correct?

- (A) Since the rate law can be expressed as **rate** = $k[A_2]^2$, tripling the concentration of A_2 will cause a 9-fold increase in the rate of the reaction.
- (B) Since the rate law can be expressed as **rate** = $k[A_2][B]$, doubling the concentrations of A_2 and B will quadruple the rate of the reaction.
- (C) Since the rate law can be expressed as **rate** = $k[A_2]^2[B]$, tripling the concentration of A_2 while keeping the concentration of B constant will triple the rate of the reaction.
- (D) Since the rate law can be expressed as **rate** = $k[A_2][B]^2$, doubling the concentration of B while keeping the concentration of A_2 constant will double the rate of the reaction.

2. $2\text{NO}(g) + \text{Cl}_2(g) \rightarrow 2\text{NOCl}(g)$

Experiment	$[\text{NO}]$	$[\text{Cl}_2]$	Initial Rate of NOCl Formation (M/s)
1	0.0250	0.0250	4.55×10^{-3}
2	0.0250	0.1000	9.10×10^{-3}
3	0.0500	0.0250	1.82×10^{-2}

The initial rates of the reaction represented by the equation shown above were measured for different initial concentrations of $\text{NO}(g)$ and $\text{Cl}_2(g)$. Based on the data given in the table above, which of the following is the rate law expression for the reaction, and why?

- (A) **Rate** = $k[\text{NO}]^2$, because the initial rate quadrupled when $[\text{NO}]$ was doubled but remained constant when $[\text{Cl}_2]$ was doubled.
- (B) **Rate** = $k[\text{NO}][\text{Cl}_2]$, because the initial rate doubled when either $[\text{NO}]$ or $[\text{Cl}_2]$ was doubled.
- (C) **Rate** = $k[\text{NO}][\text{Cl}_2]^2$, because the initial rate doubled when $[\text{NO}]$ was doubled and quadrupled when $[\text{Cl}_2]$ was doubled.
- (D) **Rate** = $k[\text{NO}]^2[\text{Cl}_2]$, because the initial rate quadrupled when $[\text{NO}]$ was doubled and doubled when $[\text{Cl}_2]$ was doubled.

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Y Pai

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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