

CHAPTER 14: COVALENT BONDING: ORBITALS

PROBLEMS WITH LEWIS STRUCTURE

- Provides no indication of the **shapes** of the molecules
- O₂ would be predicted to be diamagnetic, but it is **paramagnetic**

Paramagnetic = has unpaired electrons.



The Lewis structure of O₂ has an O=O double bond with **no unpaired electrons**, i.e. diamagnetic.

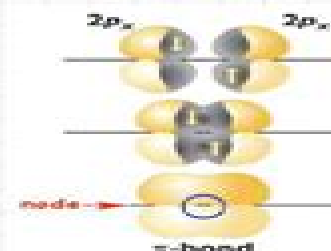
- Many compounds cannot be represented by a single structure
- Resonance structures **differ only in position of electrons** not the nuclei.
- Result in notions of hypervalent compounds and "expansion of octet"
- Can't explain molecules such as B₂H₆ (electron deficient molecules)

VALENCE-BOND THEORY "LOCALIZED ELECTRON MODEL"

- Covalent bond is created by the overlap of orbitals on two atoms

TYPES OF OVERLAP: σ and π Bonds

- Bonding in H₂**
 - Sausage shaped orbital, bond called **σ -orbital**
 - Called σ orbital because it has **no node** and is **cylindrically symmetrical**
 - Electron density is **concentrated between two nuclei**, but is still highest at the nuclei themselves.
- Bonding in N₂
 - σ orbital** is obtained by overlap of **two p_x-orbitals**
- Perpendicular p orbitals overlap in a side-by-side manner
- The resulting orbital has two lobes of electron density and is called a **π -bond**
- Called π -orbital because it **has a node** and, when viewed down the internuclear axis, it looks like a p-orbital
- π -bond has **no electron density between the nuclei along the bond axis**



General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises

Lawrence Powell Eblin



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Basic Chemistry Concepts and Exercises John Kenkel, 2011-07-08 Chemistry can be a daunting subject for the uninitiated and all too often introductory textbooks do little to make students feel at ease with the complex subject matter. Basic Chemistry Concepts and Exercises brings the wisdom of John Kenkel's more than 35 years of teaching experience to communicate the fundamentals of chemistry in a practical down-to-earth manner. Using conversational language and logically assembled graphics, the book concisely introduces each topic without overwhelming students with unnecessary detail. Example problems and end-of-chapter questions emphasize repetition of concepts, preparing students to become adept at the basics before they progress to an advanced general chemistry course. Enhanced with visualization techniques such as the first chapter's mythical microscope, the book clarifies challenging abstract ideas and stimulates curiosity into what can otherwise be an overwhelming topic. Topics discussed in this reader-friendly text include Properties and structure of matter, Atoms, molecules, and compounds, The Periodic Table, Atomic weight, formula weights, and moles, Gases and solutions, Chemical equilibrium, Acids, bases, and pH, Organic chemicals. The appendix contains answers to the homework exercises so students can check their work and receive instant feedback as to whether they have adequately grasped the concepts before moving on to the next section. Designed to help students embrace chemistry not with trepidation but with confidence, this solid preparatory text forms a firm foundation for more advanced chemistry training.

Principles of Inorganic Chemistry Brian W. Pfennig, 2015-03-03 Aimed at senior undergraduates and first-year graduate students, this book offers a principles-based approach to inorganic chemistry that, unlike other texts, uses chemical applications of group theory and molecular orbital theory throughout as an underlying framework. This highly physical approach allows students to derive the greatest benefit of topics such as molecular orbital, acid-base theory, band theory of solids, and inorganic photochemistry, to name a few. Takes a principles-based group and molecular orbital theory approach to inorganic chemistry. The first inorganic chemistry textbook to provide a thorough treatment of group theory, a topic usually relegated to only one or two chapters of texts, giving it only a cursory overview. Covers atomic and molecular term symbols, symmetry, coordinates in vibrational spectroscopy using the projection operator method, polyatomic MO theory, band theory, and Tanabe-Sugano diagrams. Includes a heavy dose of group theory in the primary inorganic textbook, most of the pedagogical benefits of integration and reinforcement of this material in the treatment of other topics such as frontier MO, acid-base theory, band theory of solids, inorganic photochemistry, the Jahn-Teller effect, and Wade's rules are fully realized. Very physical in nature compared to other textbooks in the field, taking the time to go through mathematical derivations and to compare and contrast different theories of bonding in order to allow for a more rigorous treatment of their application to molecular structure, bonding, and spectroscopy. Informal and engaging writing style, worked examples throughout the text, unanswered problems in every chapter, contains a generous use of informative, colorful illustrations.

Chemistry Therald Moeller, 2012-12-02 Chemistry with Inorganic Qualitative Analysis

is a textbook that describes the application of the principles of equilibrium represented in qualitative analysis and the properties of ions arising from the reactions of the analysis This book reviews the chemistry of inorganic substances as the science of matter the units of measure used atoms atomic structure thermochemistry nuclear chemistry molecules and ions in action This text also describes the chemical bonds the representative elements the changes of state water and the hydrosphere which also covers water pollution and water purification Water purification occurs in nature through the usual water cycle and by the action of microorganisms The air flushes dissolved gases and volatile pollutants when water seeps through the soil it filters solids as they settle in the bottom of placid lakes Microorganisms break down large organic molecules containing mostly carbon hydrogen nitrogen oxygen sulfur or phosphorus into harmless molecules and ions This text notes that natural purification occurs if the level of contaminants is not so excessive This textbook is suitable for both chemistry teachers and students

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Table of Contents General Chemistry Chapter 14 Covalent Bonding Orbitals Exericses

1. Understanding the eBook General Chemistry Chapter 14 Covalent Bonding Orbitals Exericses
 - The Rise of Digital Reading General Chemistry Chapter 14 Covalent Bonding Orbitals Exericses
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Chapter 14 Covalent Bonding Orbitals Exericses
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Chapter 14 Covalent Bonding Orbitals Exericses
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Chapter 14 Covalent Bonding Orbitals Exericses
 - Personalized Recommendations
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exericses User Reviews and Ratings

- General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises and Bestseller Lists
- 5. Accessing General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises Free and Paid eBooks
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 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises eBook Subscription Services
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises Budget-Friendly Options
- 6. Navigating General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises Compatibility with Devices
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
 - Highlighting and Note-Taking General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
 - Interactive Elements General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
- 8. Staying Engaged with General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
- 9. Balancing eBooks and Physical Books General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
 - Setting Reading Goals General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
 - Fact-Checking eBook Content of General Chemistry Chapter 14 Covalent Bonding Orbitals Exercises
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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