

(1.6-3)

From Appendix A.4, $c_p = 4.02 \text{ kJ/kg} \cdot \text{K}$ at 32.8°C
Assume $c_{p,m} = 4.02$ Note: $(t-10)^\circ\text{C} = (t-10)\text{K}$

$$q = 121\,300 \text{ kJ} = 454 \text{ kg} \left(4.02 \frac{\text{kJ}}{\text{kg} \cdot \text{K}} \right) (t - 10) \text{ K}$$

Solving, $(t-10) = 66.5 \text{ K} = 66.5^\circ\text{C}$ $t = 76.5^\circ\text{C}$

(1.6-4)

(a) A.3-9 $P = 30 \text{ psia}$ Liquid H_2O

$$\Delta H = H_{240^\circ\text{F}} - H_{40^\circ\text{F}} = 208.44 - 8.02 = \boxed{200.42 \text{ Btu/lb}_m}$$

(b) $P = 24.97 \text{ psia}$

$$H_{240^\circ\text{F}} (\text{vapor}) = 1160.7$$

$$H_{40^\circ\text{F}} (\text{liquid}) = 8.02$$

$$\Delta H = 1152.7 \text{ Btu/lb}_m$$

(c) $P = 1.0 \text{ atm abs}$

$$H_{212^\circ\text{F}} (\text{vapor}) = 1150.5$$

$$H_{40^\circ\text{F}} (\text{liquid}) = 28.1$$

$$\Delta H = -1122.4 \text{ Btu/lb}_m$$

(d) $P = 1.0 \text{ atm abs}$

$$H_{212^\circ\text{F}} (\text{vapor}) = 1150.5$$

$$H_{212^\circ\text{F}} (\text{liquid}) = 180.2$$

$$\Delta H = -970.3 \text{ Btu/lb}_m$$
$$(-2256.9 \text{ kJ/kg})$$

Geankoplis Absorption Solution Manual

Phillip C. Wankat



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