## Experiment 9 Electrochemistry I – Galvanic Cell

Introduction: Chemical reactions involving the transfer of electrons from one reactant to another are called oxidation-reduction reactions or redox reactions. In a redox reaction, two half-reactions occur, one reactant gives up electrons (undergoes oxidation) and another reactant gains electrons (undergoes reduction). A piece of zinc going into a solution as zinc ions, with each Zn atom giving up 2 electrons, is an example of an oxidation half-reaction.

$$Zn(s) \rightarrow Zn^{2+}(aq) + 2e^{-}$$
 (1)

The exidation number of Zn(s) is 0 and the exidation number of the  $Zn^{2^{n}}$  is +2. Therefore, in this half-reaction, the exidation number increases, which is another way of defining an exidation. In contrast, the reverse reaction, in which  $Zn^{2^{n}}$  ions gain 2 electrons to become Zn atoms, is an example of reduction.

$$Zn^{2+}(aq) + 2e^* \rightarrow Zn(s)$$
 (2)

In a reduction there is a decrease (or reduction) in oxidation number. Chemical equation representing half-reactions must be both mass and charge balanced. In the half-reactions above, there is one zinc on both sides of the equation. The charge is balanced because the 2+ charge on the zinc ion is balanced by two electrons, 2e', giving zero net charge on both sides.

Another example of reduction is the formation of solid copper from copper ions in solution.

$$Cu^{2+}(aq) + 2e^{\epsilon} \rightarrow Cu(s)$$
 (3)

In this half-reaction the oxidation number of the aqueous copper is +2, which decreases to 0 for the solid copper, and again charge and mass are balanced. However, no half-reaction can occur by itself. A redox reaction results when an oxidation and a reduction half-reaction are combined to complete a transfer of electrons as in the following example:

$$Zn(s) + Cu^{2+}(aq) \rightarrow Zn^{2+}(aq) + Cu(s)$$
 (4)

The electrons are not shown because they are neither reactants nor products but have simply been transferred from one species to another (from Zn to Cu<sup>2+</sup> in this case). In this redox reaction, the Zn(s) is referred to as the reducing agent because it causes the Cu<sup>2+</sup> to be reduced to Cu. The Cu<sup>2+</sup> is called the oxidizing agent because it causes the Zn(s) to be oxidized to Zn<sup>2+</sup>.

Any half-reaction can be expressed as a reduction as illustrated in the case where equation (1) can be reversed to equation (2). A measure of the tendency for a reduction to occur is its reduction potential, E, measured in units of volts. At standard conditions, 25 °C and concentrations of 1.0 M for the aqueous ions, the measured voltage of the reduction half-reaction is defined as the standard reduction potential, E°. Standard reduction potentials

# **Galvanic Cell Experiment Report**

**Michael Seilmaier** 

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#### **Table of Contents Galvanic Cell Experiment Report**

- 1. Understanding the eBook Galvanic Cell Experiment Report
  - The Rise of Digital Reading Galvanic Cell Experiment Report
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Galvanic Cell Experiment Report
  - Exploring Different Genres
  - $\circ\,$  Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Galvanic Cell Experiment Report
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Galvanic Cell Experiment Report
  - Personalized Recommendations

- Galvanic Cell Experiment Report User Reviews and Ratings
- Galvanic Cell Experiment Report and Bestseller Lists
- 5. Accessing Galvanic Cell Experiment Report Free and Paid eBooks
  - Galvanic Cell Experiment Report Public Domain eBooks
  - Galvanic Cell Experiment Report eBook Subscription Services
  - Galvanic Cell Experiment Report Budget-Friendly Options
- 6. Navigating Galvanic Cell Experiment Report eBook Formats
  - o ePub, PDF, MOBI, and More
  - Galvanic Cell Experiment Report Compatibility with Devices
  - Galvanic Cell Experiment Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - $\,\circ\,$  Adjustable Fonts and Text Sizes of Galvanic Cell Experiment Report
  - Highlighting and Note-Taking Galvanic Cell Experiment Report
  - Interactive Elements Galvanic Cell Experiment Report
- 8. Staying Engaged with Galvanic Cell Experiment Report
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Galvanic Cell Experiment Report
- 9. Balancing eBooks and Physical Books Galvanic Cell Experiment Report
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Galvanic Cell Experiment Report
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Galvanic Cell Experiment Report
  - Setting Reading Goals Galvanic Cell Experiment Report
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Galvanic Cell Experiment Report
  - Fact-Checking eBook Content of Galvanic Cell Experiment Report

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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