

4NB – VIDHATA

Date: 30/09/2014

DURATION : 2 Hrs 30 Min.

MAX MARKS : 75

INSTRUCTIONS TO THE STUDENTS :

1. All questions are compulsory.
2. Figures to the right indicate the marks.
3. Draw neat sketches / diagrams / graphs / Flow sheets whenever necessary.
4. Use of simple / scientific calculator / log tables permitted.

1. A) Answer Any Two of the following.
 - i) Predict the shape of following molecules : BeCl_2 , NH_3 , PCl_5 . 06
 - ii) a) State and explain Ionization Energy. 03
 b) What is Pajon's Rule? 03
 - iii) a) Write short note on : Covalent Bond. 03
 b) What are limitations of VSEPR theory? 03
 - iv) a) What is Electronegativity? How does it vary in period and group in the periodic table? 03
 b) Explain the Heisenberg's principle of Uncertainty? 03
- B) Answer Any Two of the following.
 - i) What is lone pair effect? How does it affect the bond angles in molecules? 04
 - ii) What are applications of VSEPR theory? 04
 - iii) Explain: BF_3 is planar whereas NH_3 is not planar. 04
 - iv) Discuss the Rutherford's atomic model. 04
2. A) Answer Any Two of the following.
 - i) a) Define the following - 03
 Hydrogen Bond, Ionic Bond, Dipole Moment.
 b) Calculate the formal charge on the atom in the following : 03
 Nitrogen in Ammonium ion
 Oxygen in Water
 Nitrogen in Nitromethane
 - ii) What are Dry Test and Wet Test? Give any two examples for each test. 06
 - iii) a) What are types of Arrows and their significance? 03
 b) Why Formic Acid is stronger than Acetic Acid? 03
 - iv) Explain the structure of following functional groups : 06
 Ether, Cyanide, Ketone.
- B) Answer Any Two of the following.
 - i) What is Inductive Effect? Explain its types. 04
 - ii) What is Semi-micro Analysis? 04
 - iii) What is Formal Charge? Explain any two examples. 04
 - iv) Explain sp^3 hybridization for Nitrogen and Oxygen. 04

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