# EXTRA PRACTICE PROBLEMS FOR MODULE #6

I. Nitric acid is manufactured with the following reaction:

How many grams of nitrogen dioxide are required to produce 5.89x103 kg HNO3 in excess water?

2. Tungsten metal, W, is produced by the following reaction:

$$WO_3 + 3H_2 \rightarrow W + 3H_2O$$

How many grams of tungsten can be made from 1.0x10° g of hydrogen in excess WO<sub>3</sub>?

3. How many grams of WO<sub>3</sub> would be used in the above reaction under the conditions specified in 82?

4. Carbon tetrachloride is made according to the following reaction:

How many grams of CS2 are needed to make 16 kg of CCL?

To make a phosphorus fertilizer (CaH<sub>4</sub>P<sub>2</sub>O<sub>8</sub>), agricultural companies use the following reaction:

$$Ca_3P_2O_8 + 2H_2SO_4 + 4H_2O \rightarrow CaH_4P_2O_8 + 2CaH_4SO_6$$

If 2.50x103 grams of Ca<sub>3</sub>P<sub>2</sub>O<sub>8</sub> are reacted with excess H<sub>2</sub>SO<sub>4</sub> and H<sub>2</sub>O, how many arms of fertilizer can be made?

6. In the following reaction:

$$Mg(s) + 2H_2O(l) \rightarrow Mg(OH)_2(s) + H_2(g)$$

What mass of Mg must you use to make 5.6x10° g of Mg(OH);?

7. Which of the following are empirical formulas?

a. SO:

b. K-S-O-

c. CsH<sub>12</sub>Os

d. PaHi

8. Decomposition of 100.0 g of a dark green powder yields 39.7 g K, 27.9 g Mn, and an unknown amount of oxygen. What is the empirical formula?

9. The mass of the molecule above is 197.1 amu. What is the molecular formula?

10. A compound has an empirical formula of CH2O. If its molecular mass is 60.0 amu, what is its molecular formula?

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J. Gvishiani

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