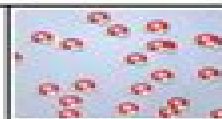


Activity B:

Temperature and molecular motion

Get the Gizmo ready:

- Click **Reset**, and select the **Micro** view.
- Set **Ice volume** to 0 cm³.
- Set **Add/remove heat energy** to 0 J/s.



Question: Why do phase changes occur?

1. **Compare:** Set the **Water temperature** to 0 °C and click **Play**. Observe the water molecules. Click **Reset**, set the **Water temperature** to 100 °C, and click **Play** again.

What do you notice?

Nothing differs, there is no difference.

2. **Observe:** Click **Reset**. The mean molecular speed of the water molecules is displayed below the container. Set the **Water temperature** to 0 °C and **Add/remove heat energy** to 400 J/s. Click **Play**.

A. How does the mean speed of the water molecules change as they are heated?

Increases the hotter water gets

B. Does the mean molecular speed change as much as the temperature as the water heats up? Explain.

Yes, as the mean molecular speed varies as temperature rises.

3. **Explain:** How is temperature related to the motions of molecules?

The molecules move more quickly as the temperature rises.

4. **Observe:** Click **Reset**. Set the **Water temperature** to 20 °C and the **Ice volume** to 50 cm³. Set **Add/remove heat energy** to 0 J/s. Click **Play**. How do the molecules in the liquid interact with the molecules in the solid?

The bonds are broken when the molecules of the liquid collide with those of the solid.

5. **Observe:** Click **Reset**. Set the **Water temperature** to 100 °C and the **Ice volume** to 50 cm³. Click **Play**. How does this situation compare to the previous one?

Ice melts more quickly

6. **Propose a theory:** Based on what you have observed, explain why you think phase changes occur. If possible, discuss your theory with your classmates and teacher.

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1. Understanding the eBook Explore Learning Gizmo Answer Key Phase Change
 - The Rise of Digital Reading Explore Learning Gizmo Answer Key Phase Change
 - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Gizmo Answer Key Phase Change
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Explore Learning Gizmo Answer Key Phase Change
 - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Gizmo Answer Key Phase Change
 - Personalized Recommendations
 - Explore Learning Gizmo Answer Key Phase Change User Reviews and Ratings
 - Explore Learning Gizmo Answer Key Phase Change and Bestseller Lists
5. Accessing Explore Learning Gizmo Answer Key Phase Change Free and Paid eBooks
 - Explore Learning Gizmo Answer Key Phase Change Public Domain eBooks
 - Explore Learning Gizmo Answer Key Phase Change eBook Subscription Services
 - Explore Learning Gizmo Answer Key Phase Change Budget-Friendly Options

6. Navigating Explore Learning Gizmo Answer Key Phase Change eBook Formats
 - ePub, PDF, MOBI, and More
 - Explore Learning Gizmo Answer Key Phase Change Compatibility with Devices
 - Explore Learning Gizmo Answer Key Phase Change Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Explore Learning Gizmo Answer Key Phase Change
 - Highlighting and Note-Taking Explore Learning Gizmo Answer Key Phase Change
 - Interactive Elements Explore Learning Gizmo Answer Key Phase Change
8. Staying Engaged with Explore Learning Gizmo Answer Key Phase Change
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Explore Learning Gizmo Answer Key Phase Change
9. Balancing eBooks and Physical Books Explore Learning Gizmo Answer Key Phase Change
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Explore Learning Gizmo Answer Key Phase Change
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Explore Learning Gizmo Answer Key Phase Change
 - Setting Reading Goals Explore Learning Gizmo Answer Key Phase Change
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Explore Learning Gizmo Answer Key Phase Change
 - Fact-Checking eBook Content of Explore Learning Gizmo Answer Key Phase Change
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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