



Name: Cassie Cantrell

Date: 04-25-2022

Student Exploration: Nuclear Decay

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: alpha particle, atomic number, beta particle, daughter product, gamma ray, isotope, mass number, nuclear decay, positron, radioactive, subatomic particle

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

The chart below gives the locations, charges, and approximate masses of three **subatomic particles**. The approximate mass of each particle is given in universal mass units (u).

Particle	Location	Charge	Approximate mass
Proton	Nucleus	1^+	1 u
Neutron	Nucleus	0	1 u
Electron	Orbitals	1^-	0 u

- The **mass number** of an atom is equal to the sum of protons and neutrons in the nucleus. A helium atom has 2 protons and 2 neutrons. What is the mass number of this atom?

4

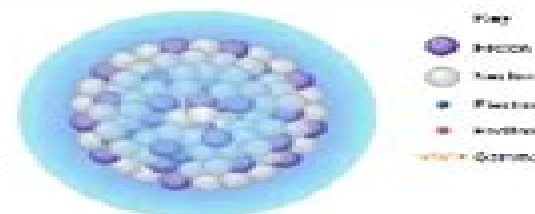
- The **atomic number** of an element is equal to the number of protons in each atom of the element. All helium atoms have 2 protons. What is the atomic number of helium?

2

Gizmo Warm-up

While most atoms are stable, some are **radioactive**, which means that they have a tendency to undergo spontaneous **nuclear decay**. The decay of radioactive atoms generally results in the emission of particles and/or energy.

Several types of nuclear decay can be explored with the *Nuclear Decay* Gizmo. On the Gizmo, check that **Alpha decay** and **Uranium** are selected.



- Click **Play** () and then click **Pause** () when the **alpha particle** is clearly visible. What is an alpha particle made of? two protons and two neutrons
- Click **Play** and observe. Besides the alpha particle, what else is emitted from the nucleus during alpha decay? a gamma ray

Gamma rays are energetic electromagnetic waves; they are often emitted in nuclear decay.

Explore Learning Nuclear Decay Gizmo Answer Key

Yijin Wang

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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Table of Contents Explore Learning Nuclear Decay Gizmo Answer Key

1. Understanding the eBook Explore Learning Nuclear Decay Gizmo Answer Key
 - The Rise of Digital Reading Explore Learning Nuclear Decay Gizmo Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Nuclear Decay Gizmo Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Explore Learning Nuclear Decay Gizmo Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Nuclear Decay Gizmo Answer Key

- Personalized Recommendations
 - Explore Learning Nuclear Decay Gizmo Answer Key User Reviews and Ratings
 - Explore Learning Nuclear Decay Gizmo Answer Key and Bestseller Lists
5. Accessing Explore Learning Nuclear Decay Gizmo Answer Key Free and Paid eBooks
 - Explore Learning Nuclear Decay Gizmo Answer Key Public Domain eBooks
 - Explore Learning Nuclear Decay Gizmo Answer Key eBook Subscription Services
 - Explore Learning Nuclear Decay Gizmo Answer Key Budget-Friendly Options
 6. Navigating Explore Learning Nuclear Decay Gizmo Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Explore Learning Nuclear Decay Gizmo Answer Key Compatibility with Devices
 - Explore Learning Nuclear Decay Gizmo Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Explore Learning Nuclear Decay Gizmo Answer Key
 - Highlighting and Note-Taking Explore Learning Nuclear Decay Gizmo Answer Key
 - Interactive Elements Explore Learning Nuclear Decay Gizmo Answer Key
 8. Staying Engaged with Explore Learning Nuclear Decay Gizmo Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Explore Learning Nuclear Decay Gizmo Answer Key
 9. Balancing eBooks and Physical Books Explore Learning Nuclear Decay Gizmo Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Explore Learning Nuclear Decay Gizmo Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Explore Learning Nuclear Decay Gizmo Answer Key
 - Setting Reading Goals Explore Learning Nuclear Decay Gizmo Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Explore Learning Nuclear Decay Gizmo Answer Key

- Fact-Checking eBook Content of Explore Learning Nuclear Decay Gizmo Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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