

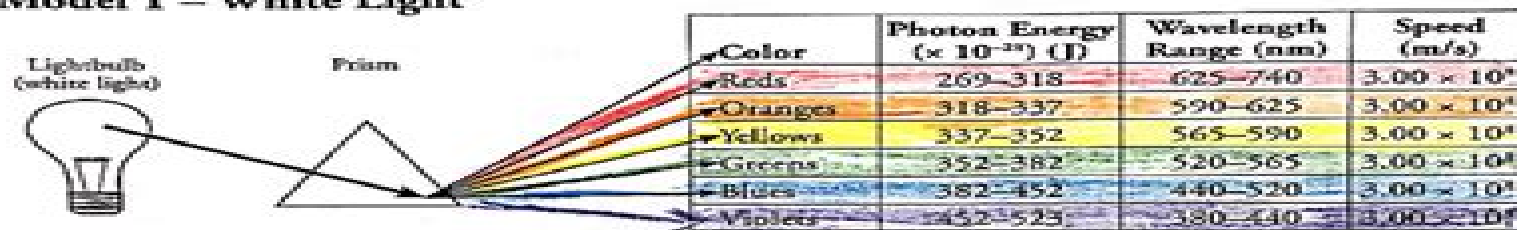
# Electron Energy and Light

How does light reveal the behavior of electrons in an atom?

## Why?

From fireworks to stars, the color of light is useful in finding out what's in matter. The emission of light by hydrogen and other atoms has played a key role in understanding the electronic structure of atoms. Trace materials, such as evidence from a crime scene, lead in paint or mercury in drinking water, can be identified by heating or burning the materials and examining the color(s) of light given off in the form of bright-line spectra.

## Model 1 – White Light



- Trace the arrows in Model 1 and shade in the table with colored pencils where appropriate.
- What happens to white light when it passes through a prism?

*The different wavelengths get separated*

- Why are the color labels in the table in Model 1 plural (i.e., "Reds" rather than "Red")?

*many different shades of a color in a range of wavelength*

- Do all colors of light travel at the same speed?

*Yes*

- Do all colors of light have the same energy? If no, which colors have the highest energy and the least energy, respectively?

*All colors do not have the same energy. Blues + violets highest  
Reds + oranges lowest*

- Consider the light illustrated in Model 1.

- Which color corresponds to the longest wavelengths?

*Reds*

- Which color corresponds to the shortest wavelengths?

*Violets*

- Write a sentence that describes the relationship between wavelength and energy of light.

*As the wavelength gets shorter the energy of light increases*

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**JL Elias**



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As light can show wave particle duality so can an electron beam Its measured wavelength is about the same as X rays This means electrons can be diffracted The famous double slit experiment where a single electron appears to interfere with itself cannot be explained by classical physics and so we enter the strange world of quantum mechanics The birth of the quantum computer is not far away and will be much faster than existing computers Finally all chemical reactions are the result of electron movement between reactants      *The Energy of Photo-electrons from Sodium and Potassium as a Function of the Frequency of the Incident Light ...* William Henry Kadesch,1914

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