

## KEY

## Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

## Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as  $\text{NaCl}$ . In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

### Model 1 – Ion Charges for Selected Elements

1	H <sup>+</sup>									
2	Li <sup>+</sup>	Be <sup>2+</sup>						N <sup>3-</sup>	O <sup>2-</sup>	F <sup>-</sup>
3	Na <sup>+</sup>	Mg <sup>2+</sup>	Transition elements				Al <sup>3+</sup>		P <sup>3-</sup>	S <sup>2-</sup>
4	K <sup>+</sup>	Ca <sup>2+</sup>	Fe <sup>2+</sup> Fe <sup>3+</sup>	Ni <sup>2+</sup> Ni <sup>3+</sup>	Cu <sup>+</sup> Cu <sup>2+</sup>	Zn <sup>2+</sup>				Br <sup>1-</sup>
5	Rb <sup>+</sup>	Sr <sup>2+</sup>			Ag <sup>1+</sup>			Sn <sup>2+</sup> Sn <sup>4+</sup>		I <sup>-</sup>
6		Ba <sup>2+</sup>				Hg <sup>2+</sup> Hg <sup>1+</sup>		Pb <sup>2+</sup> Pb <sup>4+</sup>		
	Cations						Anions			

1. Based on the information in Model 1:
- Identify three elements that form only one cation.  
 $H^+$   $Li^+$   $Na^+$   $K^+$   $Rb^+$   $Be^{+2}$   $Mg^{+2}$   $Ca^{+2}$   $Sr^{+2}$   $Ba^{+2}$   $Zn^{+2}$   $Ag^+$   $Al^{+3}$
  - Identify three elements that form only one anion.  
 $N^{-3}$   $P^{-3}$   $O^{-2}$   $S^{-2}$   $F^-$   $Cl^-$   $Br^-$   $I^-$
  - Identify three elements that form more than one cation.  
 $Fe$   $Ni$   $Cu$   $Hg$   $Sn$   $Pb$
  - In what region of the periodic table are these "multiple ion" elements usually located?  
**Transition Metals**

2. Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(x) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



3. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



# Naming Molecular Compounds Pogil Answers Key

**Bull, Prince Hycy, Patterson, Gerrelyn  
Chunn**



## **Naming Molecular Compounds Pogil Answers Key:**

Redefining Teacher Education and Teacher Preparation Programs in the Post-COVID-19 Era Bull, Prince Hycy, Patterson, Gerrelyn Chunn, 2021-12-17 Due to the COVID 19 pandemic teacher preparation programs modified their practices to fit the delivery modes of school districts while developing new ways to prepare candidates Governmental agencies established new guidelines to fit the drastic shift in education caused by the pandemic and P 12 school systems made accommodations to support teacher education candidates The pandemic disrupted all established systems and norms however many practices and strategies emerged in educator preparation programs that will have a lasting positive impact on P 20 education and teacher education practices Such practices include the reevaluation of schooling practices with shifts in engagement strategies instructional approaches technology utilization and supporting students and their families Redefining Teacher Education and Teacher Preparation Programs in the Post COVID 19 Era provides relevant innovative practices implemented across teacher education programs and P 20 settings including delivery models training procedures theoretical frameworks district policies and guidelines state national and international standards digital design and delivery of content and the latest empirical research findings on the state of teacher education preparation The book showcases best practices used to shape and redefine teacher education through the COVID 19 pandemic Covering topics such as online teaching practices simulated teaching experiences and emotional learning this text is essential for preservice professionals paraprofessionals administrators P 12 faculty education preparation program designers principals superintendents researchers students and academicians

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