



Name:

Date:

Student Exploration: Ionic Bonds

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: chemical family, electron affinity, ion, ionic bond, metal, nonmetal, octet rule, shell, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. Nate and Clara are drawing pictures with markers. There are 8 markers in a set. Nate has 9 markers and Clara has 7. What can Nate and Clara do so that each of them has a full set?

Nate can give Clara a marker.

2. Maggie is sitting at a table with Fred and Florence. Maggie has 10 markers, but Fred and Florence each have only 7 markers. How can they share markers so each has 8?

Maggie can give Fred and Florence each a marker.

Gizmo Warm-up

Just like students sharing markers, atoms sometimes share or swap electrons. By doing this, atoms form bonds. The *Ionic Bonds* Gizmo allows you to explore how **ionic bonds** form.

To begin, check that **Sodium (Na)** and **Chlorine (Cl)** are selected from the menus at right. Click **Play** () to see electrons orbiting the nucleus of each atom. (Note: These atom models are simplified and not meant to be realistic.)



1. Each atom consists of a central nucleus and several **shells** that contain electrons. The outermost electrons are called **valence electrons**. (Inner electrons are not shown.)

How many valence electrons does each atom have?

Sodium:

Chlorine:

2. Click **Pause** (). Elements can be classified as **metals** and **nonmetals**. Metals do not hold on to their valence electrons very tightly, while nonmetals hold their electrons tightly. **Electron affinity** is a measure of how tightly the valence electrons are held.

- A. Try pulling an electron away from each atom. Based on this experiment, which atom is a metal?

Which is a nonmetal?

- B. Try moving an electron from the metal to the nonmetal. What happens?

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Table of Contents Ionic Bonds Gizmo Activity A Ions Key

1. Understanding the eBook Ionic Bonds Gizmo Activity A Ions Key
 - The Rise of Digital Reading Ionic Bonds Gizmo Activity A Ions Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonds Gizmo Activity A Ions Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonds Gizmo Activity A Ions Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonds Gizmo Activity A Ions Key

- Personalized Recommendations
 - Ionic Bonds Gizmo Activity A Ions Key User Reviews and Ratings
 - Ionic Bonds Gizmo Activity A Ions Key and Bestseller Lists
5. Accessing Ionic Bonds Gizmo Activity A Ions Key Free and Paid eBooks
 - Ionic Bonds Gizmo Activity A Ions Key Public Domain eBooks
 - Ionic Bonds Gizmo Activity A Ions Key eBook Subscription Services
 - Ionic Bonds Gizmo Activity A Ions Key Budget-Friendly Options
 6. Navigating Ionic Bonds Gizmo Activity A Ions Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Bonds Gizmo Activity A Ions Key Compatibility with Devices
 - Ionic Bonds Gizmo Activity A Ions Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonds Gizmo Activity A Ions Key
 - Highlighting and Note-Taking Ionic Bonds Gizmo Activity A Ions Key
 - Interactive Elements Ionic Bonds Gizmo Activity A Ions Key
 8. Staying Engaged with Ionic Bonds Gizmo Activity A Ions Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonds Gizmo Activity A Ions Key
 9. Balancing eBooks and Physical Books Ionic Bonds Gizmo Activity A Ions Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonds Gizmo Activity A Ions Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ionic Bonds Gizmo Activity A Ions Key
 - Setting Reading Goals Ionic Bonds Gizmo Activity A Ions Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ionic Bonds Gizmo Activity A Ions Key

- Fact-Checking eBook Content of Ionic Bonds Gizmo Activity A Ions Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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