

## Questions

2. Following the rules for stabilisation for the last two questions, write correct chemical formulae for each of the following compounds:

Compound	Formula
(a) Lithium chloride	$\text{Li}^{+} / \text{Cl}^{-} \rightarrow \text{LiCl}$
(b) Magnesium nitride	$\text{Mg}^{2+} / \text{N}^{3-} \rightarrow \text{Mg}_3\text{N}_2$
(c) Strontium sulfide	$\text{Sr}^{2+} / \text{S}^{2-} \rightarrow \text{SrS}$
(d) Rubidium fluoride	$\text{Rb}^{+} / \text{F}^{-} \rightarrow \text{RbF}$
(e) Lithium oxide	$\text{Li}^{+} / \text{O}^{2-} \rightarrow \text{Li}_2\text{O}$
(f) Sodium selenide	$\text{Na}^{+} / \text{Se}^{2-} \rightarrow \text{Na}_2\text{Se}$
(g) Potassium chloride	$\text{K}^{+} / \text{Cl}^{-} \rightarrow \text{KCl}$
(h) Calcium phosphide	$\text{Ca}^{2+} / \text{P}^{3-} \rightarrow \text{Ca}_3\text{P}_2$
(i) Barium oxide	$\text{Ba}^{2+} / \text{O}^{2-} \rightarrow \text{BaO}$
(j) Aluminium nitride	$\text{Al}^{3+} / \text{N}^{3-} \rightarrow \text{AlN}$

3. Use your answers to the Key Questions and the Exercise Questions to draw a conclusion about the order of ions in each compound of the elements in the compounds you drew the most groups (strontium, rubidium and aluminium nitride, lithium chloride and potassium chloride).

# Ionic Bonding Chemistry If8766 Answers

**G Thomas**



## **Ionic Bonding Chemistry If8766 Answers:**

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