

II PUC CHEMISTRY SYLLABUS BLOW-UP

UNIT-I

Solid State

8 hrs

General characteristics of solids: amorphous and crystalline solids – examples, differences. Classification of crystalline solids – based on binding forces: molecular solids – (non-polar, polar, H-bonded), ionic solids, metallic solids, covalent or network solids – examples for all.

Definitions – lattice point, crystal lattice, unit cell, coordination number. Parameters of a unit cell, names of seven crystal systems, calculation of number of atoms in a cubic unit cell – simple cubic, fcc, bcc. Close packing in two-dimensional and three dimensional lattices – brief information, voids – types of voids, tetrahedral and octahedral and their relative numbers, calculation of the formula of the compounds based on the number of voids filled. Packing in solids – calculation of packing efficiency- fcc/bcc, bcc, simple cubic. Formula to calculate the density of the unit cell to be assumed- use the formula to calculate a , d , z , M , N_A . Numerical problems.

Point defects-types, a brief account of Frenkel and Schottky defects, metal excess defect and metal deficiency defect with examples.

Electrical properties: classification into conductors, insulators and semiconductors – their comparison based on band theory of metals, n-type and p-type semiconductors – differences and examples. Magnetic properties of substances – paramagnetism, diamagnetism and ferromagnetism, examples.

UNIT-II

Solutions

9 hrs

Types of Solutions – binary – gaseous, liquid and solids, expressing the concentration of a solution of a solid in a liquid – mole fraction, molarity and molality. Solubility, solubility of a gas in a liquid – Henry's law, graph of partial pressure of a gas vs its mole fraction in solution, effect of pressure, temperature, applications of Henry's law.

Solution of liquid in liquid – Raoult's law- statement, mathematical expression, numerical problems, ideal solution – characteristics, graph, non-ideal solution –types – their characteristics and differences, examples, azeotropes – types, examples.

Solution of a solid in a liquid – Raoult's law – colligative properties – relative lowering of vapour pressure, elevation in boiling point, depression in freezing point, graphs for elevation in boiling point and depression in freezing point, SI units for K_b , K_f , osmosis – osmotic pressure, isotonic, hypertonic, hypotonic solutions, reverse osmosis – application in desalination of water. Numerical problems on determination of molar mass using colligative properties.

Abnormal molar mass, van't Hoff factor i , value of i for non-electrolytes and solutes that associate or dissociate in dilute solution.

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Yi-Tong Ma



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Rules, Regulations and Specifications Public Utilities Commission of Ohio,1929 Federal Power Commission Reports United States. Federal Power Commission,1961 Contains all the formal opinions and accompanying orders of the Federal Power Commission In addition to the formal opinions there have been included intermediate decisions which have become final and selected orders of the Commission issued during such period *Nuclear Waste Disposal* United States. Congress. Senate. Committee on Environment and Public Works. Subcommittee on Nuclear Regulation,1980

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