Section 2-1 The Nature of Matter (pages 35-39)

This section identifies the three particles that make up atoms. It also explains how atoms of the same element can have a different number of neutrons and describes the two main types of chemical bonds.

Atoms (page 35)

- 1. The basic unit of matter is called a(an) .
- 2. Describe the nucleus of an atom.
- 3. Complete the table about subatomic particles.

SUBATOMIC PARTICLES

Particle	Charge	Location in Atom
	Positive	
	Neutral	7
	Negative	

4. Why are atoms neutral despite having charged particles?

Elements (page 36)

- 5. What is a chemical element?
- 6. What does an element's atomic number represent?

Chemical Compounds (page 37)

- 7. What is a chemical compound?
- 8. What are the elements in salt? How many of each element are there?

Chemical Bonds (pages 38-39)

- 9. What holds atoms in compounds together?
- 10. Complete the table about the main types of chemical bonds.

CHEMICAL BONDS

Chemical bonds

Type Formed When		
Covalent bond		
Ionic bond		

- 11. What is an ion?
- 12. Is the following sentence true or false?

An atom that loses electrons has a negative charge.

13. The structure that results when atoms are joined together by covalent bonds is called a(an).

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E Durkheim

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