

## Notes Lewis Dot Structures

Representative of only the valence electrons

Valence electrons: electrons in the outermost energy level.

Periodic Table Label:

12345678

Ex: Na



Ex: N



Ex: S



## Notes Bonding

Chemical Formula: shows the type and number of atoms in the smallest representative unit of a substance

Subscript	Superscript
$\text{Cl}_2$ Tells the number of elements present: 2 Cls	$\text{Cl}^{1-}$ Tells the charge of the ion - means a GAIN of electrons - anion
$\text{H}_2\text{O}$ H: 2 O: 1	$\text{Mg}^{2+}$ + Means a LOSS of electrons - cation
$\text{H}_3\text{PO}_4$ H: 3 P: 1 O: 4	$\text{CO}_3^{2-}$ A charge on multiple elements means it is a polyatomic ion; the charge is shared over the entire compound

# Guided Notes 8 1 The Covalent Bond Answer Key

**ML Morrison**



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