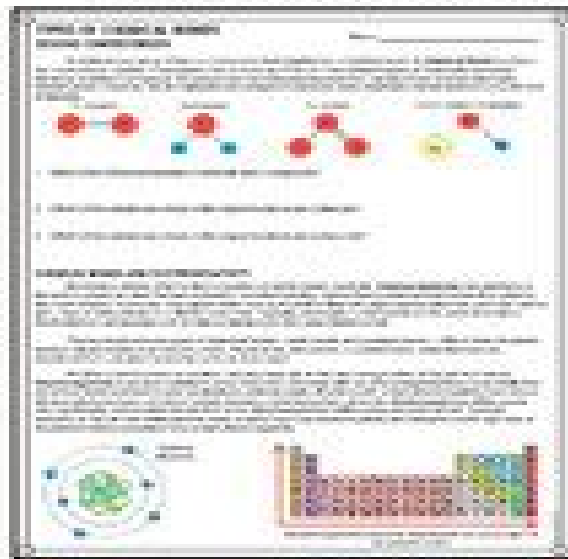


# GUIDED READING WITH CHUNKED QUESTIONS THROUGHOUT & POST-READING QUESTIONS



Types of Chemical Bonds

1. Ionic bonds are formed between a metal and a non-metal. The metal atom loses one or more electrons to become a positive ion, and the non-metal atom gains one or more electrons to become a negative ion. The resulting ions are attracted to each other by strong electrostatic forces, forming an ionic bond.

2. Covalent bonds are formed between two non-metal atoms. The atoms share one or more pairs of electrons, forming a covalent bond.

3. Metallic bonds are formed between metal atoms. The metal atoms are arranged in a regular lattice, and the outer electrons are delocalized, forming a 'sea of electrons' that holds the metal ions together.

Diagram showing the formation of ionic, covalent, and metallic bonds.



Properties of Ionic Compounds

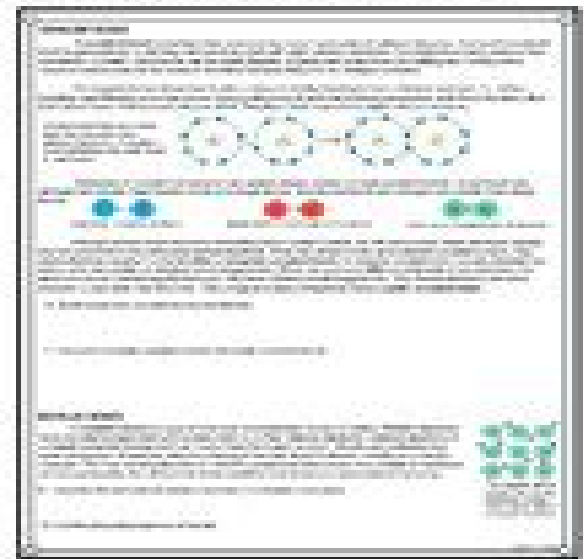
1. High melting and boiling points due to strong electrostatic forces.

2. Soluble in water, forming electrolyte solutions.

3. Conduct electricity when molten or in aqueous solution.

4. Brittle solids.

Diagram showing the structure of an ionic compound lattice.



Properties of Covalent Compounds

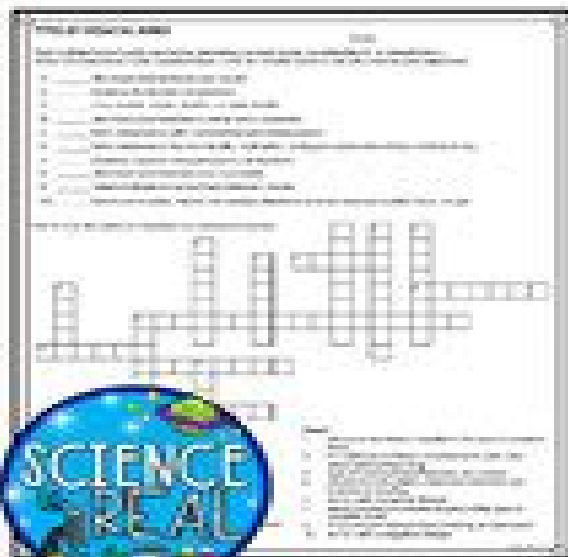
1. Low melting and boiling points due to weak intermolecular forces.

2. Insoluble in water.

3. Do not conduct electricity.

4. Can be gases, liquids, or solids.

Diagram showing the structure of a covalent molecule.



Properties of Metallic Compounds

1. High melting and boiling points due to strong metallic bonds.

2. Soluble in water, forming electrolyte solutions.

3. Conduct electricity.

4. Malleable and ductile.

Diagram showing the structure of a metallic lattice.

- ✓ IONIC, COVALENT, & METALLIC BONDS
- ✓ 13 CHUNKED QUESTIONS IN READING
- ✓ 23 POST-READING QUESTIONS
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