

# NAMING IONIC COMPOUNDS

Name \_\_\_\_\_

Name the following compounds using the Stock Naming System.

or traditional  
(where appropriate)

1.  $\text{CaCO}_3$  calcium carbonate
2.  $\text{KCl}$  potassium chloride
3.  $\overset{+2}{\text{Fe}}\text{SO}_4$  iron(II) sulfate ferrous sulfate
4.  $\text{LiBr}$  lithium bromide
5.  $\text{MgCl}_2$  magnesium chloride
6.  $\overset{+3}{\text{Fe}}\text{Cl}_3$  iron(III) chloride ferric chloride
7.  $\text{Zn}_3(\text{PO}_4)_2$  zinc phosphate
8.  $\text{NH}_4\text{NO}_3$  ammonium nitrate
9.  $\text{Al}(\text{OH})_3$  aluminum hydroxide
10.  $\overset{+1}{\text{Cu}}\text{C}_2\text{H}_3\text{O}_2$  copper(I) acetate cuprous acetate
11.  $\overset{+2}{\text{Pb}}\text{SO}_3$  lead(II) sulfite plumbous sulfite
12.  $\text{NaClO}_3$  sodium chlorate
13.  $\text{CaC}_2\text{O}_4$  calcium oxalate
14.  $\overset{+3}{\text{Fe}}_2\text{O}_3$  iron(III) oxide ferric oxide
15.  $(\text{NH}_4)_3\text{PO}_4$  ammonium phosphate
16.  $\text{NaHSO}_4$  sodium bisulfate / sodium hydrogen sulfate
17.  $\text{Hg}_2\text{Cl}_2$  mercury(I) chloride mercurous chloride
18.  $\text{Mg}(\text{NO}_2)_2$  magnesium nitrite
19.  $\overset{+2}{\text{Cu}}\text{SO}_4$  copper(II) sulfate cupric sulfate
20.  $\text{NaHCO}_3$  sodium bicarbonate / sodium hydrogen carbonate
21.  $\overset{+3}{\text{Ni}}\text{Br}_3$  nickel(III) bromide
22.  $\text{Be}(\text{NO}_3)_2$  beryllium nitrate
23.  $\text{ZnSO}_4$  zinc sulfate
24.  $\overset{+3}{\text{Au}}\text{Cl}_3$  gold(III) chloride
25.  $\text{KMnO}_4$  potassium permanganate

# Naming Compounds Containing Polyatomic Ions

## Answers

**L Cohen**



## **Naming Compounds Containing Polyatomic Ions Answers:**

The Practice of Chemistry Donald J. Wink, Sharon Fetzner-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math. The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving. The book's unique focus on actual chemical practice, extensive study tools, and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors. This special PowerPoint tour of the text was created by Don Wink. [http://www.bfwpub.com/pdfs/wink/POCPowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt) 832KB

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chemistry This study guide will be of great value to chemistry teachers and students

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