

**Assignment #2 – Compound Names and Formulas**  
**Multi-valent ions only**

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

1. copper (II) oxide CuO
2. mercury (I) oxide Hg<sub>2</sub>O
3. gold (III) chloride AuCl<sub>3</sub>
4. thallium (III) bromide TlBr<sub>3</sub>
5. bismuth (V) oxide Bi<sub>2</sub>O<sub>5</sub>
6. terbium (III) oxide Tb<sub>2</sub>O<sub>3</sub>
7. uranium (VI) oxide UO<sub>3</sub>
8. protactinium (IV) bromide PaBr<sub>4</sub>
9. cerium (III) oxide Ce<sub>2</sub>O<sub>3</sub>
10. arsenic (V) sulphide As<sub>2</sub>S<sub>5</sub>

11. manganese (III) oxide Mn<sub>2</sub>O<sub>3</sub>
12. vanadium (II) bromide VBr<sub>2</sub>
13. niobium(V) oxide Nb<sub>2</sub>O<sub>5</sub>
14. titanium (III) oxide Ti<sub>2</sub>O<sub>3</sub>
15. titanium (III) nitride TiN
16. iron (II) oxide FeO
17. cobalt (II) phosphide Co<sub>3</sub>P<sub>2</sub>
18. tin (II) oxide SnO
19. thulium (II) bromide TmBr<sub>2</sub>
20. copper (II) bromide CuBr<sub>2</sub>

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

1. SnCl<sub>4</sub> Tin (IV) Chloride
2. BiBr<sub>5</sub> Bismuth (V) Bromide
3. PoO<sub>2</sub> Polonium (IV) Oxide
4. PbI<sub>2</sub> Lead (II) Iodide
5. HgO Mercury (II) Oxide
6. HgCl Mercury (I) Chloride
7. Au<sub>2</sub>O<sub>3</sub> Gold (III) Oxide
8. FeCl<sub>2</sub> Iron (II) Chloride

9. PdF<sub>4</sub> Palladium (IV) Fluoride
10. Os<sub>2</sub>O<sub>3</sub> Osmium (III) Oxide
11. MoBr<sub>2</sub> Molybdenum (II) Bromide
12. VCl<sub>5</sub> Vanadium (V) Chloride
13. Mn<sub>2</sub>O<sub>3</sub> Manganese (III) Oxide
14. CoO Cobalt (II) Oxide
15. Np<sub>2</sub>O<sub>3</sub> Neptunium (III) Oxide
16. V<sub>2</sub>O<sub>5</sub> Vanadium (V) Oxide

# Naming Balancing And Formula Writing Answer Key

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## **Table of Contents Naming Balancing And Formula Writing Answer Key**

1. Understanding the eBook Naming Balancing And Formula Writing Answer Key
  - The Rise of Digital Reading Naming Balancing And Formula Writing Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Naming Balancing And Formula Writing Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Naming Balancing And Formula Writing Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Naming Balancing And Formula Writing Answer Key
  - Personalized Recommendations
  - Naming Balancing And Formula Writing Answer Key User Reviews and Ratings
  - Naming Balancing And Formula Writing Answer Key and Bestseller Lists
5. Accessing Naming Balancing And Formula Writing Answer Key Free and Paid eBooks
  - Naming Balancing And Formula Writing Answer Key Public Domain eBooks
  - Naming Balancing And Formula Writing Answer Key eBook Subscription Services
  - Naming Balancing And Formula Writing Answer Key Budget-Friendly Options
6. Navigating Naming Balancing And Formula Writing Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Naming Balancing And Formula Writing Answer Key Compatibility with Devices
  - Naming Balancing And Formula Writing Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Naming Balancing And Formula Writing Answer Key
  - Highlighting and Note-Taking Naming Balancing And Formula Writing Answer Key
  - Interactive Elements Naming Balancing And Formula Writing Answer Key
8. Staying Engaged with Naming Balancing And Formula Writing Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Naming Balancing And Formula Writing Answer Key
- 9. Balancing eBooks and Physical Books Naming Balancing And Formula Writing Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Naming Balancing And Formula Writing Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Naming Balancing And Formula Writing Answer Key
  - Setting Reading Goals Naming Balancing And Formula Writing Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Naming Balancing And Formula Writing Answer Key
  - Fact-Checking eBook Content of Naming Balancing And Formula Writing Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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