

NACE CP3 - Chpt 1. Mechanisms of Corrosion Exam study guides Questions With 100% Correct Answers (2024 / 2025) (Verified by Experts)

Corrosion Zone - ANSWER $\text{Fe}^0 \rightarrow \text{Fe}^{++} + 2\text{e}^-$ (soluble - reaction can keep occurring)

Immunity to Passive - ANSWER $\text{Fe}^0 + 2\text{H}_2\text{O} \rightarrow \text{Fe}(\text{OH})_3 + 2\text{H}^+ + 2\text{e}^-$ (insoluble - builds up at the interface)

Corrosion to Passive - ANSWER $\text{Fe}^{++} + 3\text{H}_2\text{O} \rightarrow \text{Fe}(\text{OH})_3 + 3\text{H}^+ + 1\text{e}^-$

More Active (+ or -) - ANSWER Electronegative - Less Noble

Less Active (+ or -) - ANSWER Electropositive - More Noble

Nerst Equation - ANSWER A metal electrode potential is a function of the metal ion activity, which is related to metal ion concentration.

As metal ion concentration increases (M^{n+}) the potential becomes more electropositive.

At 1 molar solution of metal ions.

$E_m = \dots\dots\dots \ln(a^{\text{Mn+}}/a^{\text{Mo}})$

Cathode Reaction - ANSWER Reduction Reaction

Anode Reaction - ANSWER Oxidation Reaction

SHE to CSE - ANSWER +316 mV

SHE to SSC (SJ) - ANSWER +256 mV

SHE to ZRE - ANSWER -800 mV

SSC to CSE - ANSWER +60 mV

In non-aerated soils and waters the typical reduction reaction is ... - **ANSWER**

Reduction of Hydrogen

$\text{H}^+ + \text{e}^- \rightarrow \text{H}_0$

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