

Ionic and Covalent Bonds-Study Guide

We will be defining the following terms:

- | | | |
|-------------------|-----------------|--------------------|
| • ionic bond | • anion | • polyatomic ion |
| • Lewis structure | • cation | • hydrogen bonding |
| • molecule | • chemical bond | • valence electron |
| • polar bond | • ionic bond | |
| • non-polar bond | • ion | |

1. A group of covalently bonded atoms that acts together as one charged atom is a polyatomic ion.
2. A chemical bond that occurs when atoms share electrons is called a covalent bond.
3. The octet rule tells you how many electrons an atom must gain, lose, or share to become stable.
4. The oxidation number of an atom is shown with a roman numeral in parenthesis.
5. How many electrons are needed to fill the outer energy levels of most atoms for the atom to be chemically stable? 8 (Eight)
6. Why do the noble gases NOT form compounds readily? They have 8 electrons in their outer electron shell.
7. What is the total number of atoms in the compound $\text{Ca}(\text{OH})_2$? 4 (four)
8. Explain why it is correct to say that a compound like calcium chloride, CaCl_2 , contains both ionic and covalent bonds. Chlorine is a polyatomic ion so when bonded and there is ionic bond with Calcium.
9. Name or write the formula for the following ionic compounds:

a. <u>Al_2O_3 Aluminum Oxide</u>	d. <u>magnesium (II) oxide</u>
b. <u>NH_4Cl Ammonium Chloride</u>	e. <u>Mg_2O_3</u>
c. <u>FeCl_2 Iron (II) Chloride</u>	f. <u>lead (IV) nitrate</u>
	<u>$\text{Pb}(\text{NO}_3)_4$</u>

Ions Chemthink Answers

Brendan G. Carr

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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Ions in Solution (1) (sound Recording). John Burgess,1984 Reactions of Ions William Edward Marsden,1996

Ions in Solution Ronald Wilfrid Gurney,1962 **Complex Ions** Arthur Jaques,2015-07-07 Excerpt from Complex Ions

In Aqueous Solutions In compiling this volume the needs and criticism of a large class of students unversed in physical chemistry have been especially kept in view and it is considered that the introduction of some elementary matter such as proofs of formulae which the advanced reader will not require is by no means out of place In giving an account of the methods in Chapters III VI it was found necessary to introduce examples but these were made as brief as possible in order to avoid confusing these chapters with the later ones which deal with practical investigations where more than one method is generally used at a time The tension experiments in Chapter VIII form a method of investigation in which the examination of different salts shows so little variation that it appeared unnecessary to devote a separate chapter to the method The chief aim of the book is to give some account of the more important experimental work in this subject and no apology is offered for the absence of theories of valency Chapter X contains an account of some results besides the identification of complex compounds which have been arrived at by similar methods and which are likely to form the basis of further experiments About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books Find more at www.forgottenbooks.com This book is a reproduction of an important historical work Forgotten Books uses state of the art technology to digitally reconstruct the work preserving the original format whilst repairing imperfections present in the aged copy In rare cases an imperfection in the original such as a blemish or missing page may be replicated in our edition We do however repair the vast majority of imperfections successfully any imperfections that remain are intentionally left to preserve the state of such historical works *Ions in Solution* ,1972 **Ions in Solution** Ronald Wilfrid Gurney,1936

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