

# IONIC AND METALLIC BONDING

## Homework

### Set 1: IONS

- For each element below, state (i) the number of valence electrons in the atom, (ii) the electron dot structure, and (iii) the chemical symbol(s) for the most stable ion.
  - Ba
  - I
  - K
- How many valence electrons does each of the following atoms have?
  - gallium
  - fluorine
  - selenium
- Write the electron configuration for each of the following atoms and ions.
  - Ca
  - chlorine atom
  - $\text{Na}^+$
  - phosphide ion
  - $\text{O}^{2-}$
- What is the relationship between the group number of the representative elements and the number of valence electrons?
- How many electrons will each element gain or lose in forming an ion? State whether the resulting ion is a cation or an anion.
  - strontium
  - aluminum
  - tellurium
  - rubidium
  - bromine
  - phosphorus
- Give the name and symbol of the ion formed when
  - a chlorine atom gains one electron.
  - a potassium atom loses one electron.
  - an oxygen atom gains two electrons.
  - a barium atom loses two electrons.
- How many electrons are lost or gained in forming each of the following ions?
  - $\text{Mg}^{2+}$
  - $\text{Br}^-$
  - $\text{Ag}^+$
  - $\text{Fe}^{3+}$
- Classify each of the following as a cation or an anion.
  - $\text{Na}^+$
  - $\text{Cu}^{2+}$
  - $\text{I}^-$
  - $\text{O}^{2-}$
  - $\text{Ca}^{2+}$
  - $\text{Cs}^+$

### Set 2: IONIC BONDS AND IONIC COMPOUNDS

- Use electron dot structures to predict the formula of the ionic compounds formed when the following elements combine.
  - sodium and bromine
  - sodium and sulfur
  - calcium and iodine
  - aluminum and oxygen
  - barium and chlorine
- Which of these combinations of elements are most likely to react to form ionic compounds?
  - sodium and magnesium
  - barium and sulfur
  - potassium and iodine
  - oxygen and argon
- What is the meaning of coordination number?
- How is the coordination number determined?

# Ionic And Metallic Bonding Test A Answers

**J Rink**



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